

REVIEW FOR THERMOCHEMISTRY TEST

1. The combustion of methanol is shown by the following equation:



- a) Using the data below, find the heat of reaction for the combustion of methanol.



- b) What is the molar enthalpy of combustion for methanol?

$$\Delta H = -1278 \text{ kJ/2 mol} = -639 \text{ kJ/mol}$$

- c) Express the energy changes for the target reaction in the form $\Delta H_r =$

$$\Delta H_{\text{combustion CH}_3\text{OH}} = -639 \text{ kJ/mol}$$

- d) Draw the potential energy diagram to illustrate the energy changes for the target reaction.



Thermochemistry Test A

Huangqi Zhang



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