Thermochemistry Test Preview

Matching

Match each item with the correct statement below.

a. calorimeter
b. calorie
c. joule
d. enthalpy
e. specific heat
f. heat capacity

_____ 1. quantity of heat needed to raise the temperature of 1 g of water by 1°C
_____ 2. SI unit of energy
_____ 3. quantity of heat needed to change the temperature of 1 g of a substance by 1°C
_____ 4. quantity of heat needed to change the temperature of an object by 1°C
_____ 5. device used to measure the heat absorbed or released during a chemical or physical process
_____ 6. heat content of a system at constant pressure

Match each item with the correct statement below.

a. heat of reaction
b. heat of formation
c. Hess's law of heat summation
d. heat of fusion
e. heat of solution

_____ 7. the enthalpy change for a chemical reaction exactly as it is written
_____ 8. the enthalpy change caused by dissolving a substance
_____ 9. the energy required to melt a solid at its melting point
_____ 10. the change in enthalpy that accompanies the formation of a compound from its elements
_____ 11. states that if you add two or more thermochemical equations to give a final equation, you can also add the heats of reaction to give the final heat of reaction

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

_____ 12. What happens to the energy produced by burning gasoline in a car engine?
a. The energy is lost as heat in the exhaust.
b. The energy is transformed into work to move the car.
c. The energy heats the parts of the engine.
d. all of the above

_____ 13. A piece of metal is heated, then submerged in cool water. Which statement below describes what happens?
a. The temperature of the metal will increase.
b. The temperature of the water will increase.
c. The temperature of the water will decrease.
d. The temperature of the water will increase and the temperature of the metal will decrease.

_____ 14. How does a calorie compare to a joule?
a. A calorie is smaller than a joule.
b. A calorie is larger than a joule.
c. A calorie is equal to a joule.
d. The relationship cannot be determined.

_____ 15. What would likely happen if you were to touch the flask in which an endothermic reaction were occurring?
a. The flask would probably feel cooler than before the reaction started.
b. The flask would probably feel warmer than before the reaction started.
c. The flask would feel the same as before the reaction started.
d. none of the above

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