

Annotated periodic table of the elements

1	H 1.008 Hydrogen																		18	Ar 39.95 Argon			
2	Li 6.94 Lithium		Be 9.012 Beryllium																		19	K 39.10 Potassium	
3	Na 22.99 Sodium		Mg 24.31 Magnesium																		20	Ca 40.08 Calcium	
4	K 39.10 Potassium		Ca 40.08 Calcium																		21	Sc 44.96 Scandium	
5	Rb 85.47 Rubidium		Sr 87.62 Strontium																		22	Ti 47.87 Titanium	
6	Cs 132.91 Cesium		Ba 137.33 Barium																		23	V 50.94 Vanadium	
7	Fr 223 Francium		Ra 226 Radium																		24	Cr 52.00 Chromium	
8																					25	Mn 54.94 Manganese	
9																					26	Fe 55.85 Iron	
10																					27	Co 58.93 Cobalt	
11																					28	Ni 58.69 Nickel	
12																					29	Cu 63.55 Copper	
13																					30	Zn 65.38 Zinc	
14																					31	Ga 69.72 Gallium	
15																					32	Ge 72.64 Germanium	
16																					33	As 74.92 Arsenic	
17																					34	Se 78.96 Selenium	
18																					35	Br 79.90 Bromine	
19																					36	Kr 83.80 Krypton	
20																					37	Rb 85.47 Rubidium	
21																					38	Sr 87.62 Strontium	
22																					39	Y 88.91 Yttrium	
23																					40	Zr 91.22 Zirconium	
24																					41	Nb 92.91 Niobium	
25																					42	Mo 95.94 Molybdenum	
26																					43	Tc 98.91 Technetium	
27																					44	Ru 101.07 Ruthenium	
28																					45	Rh 101.07 Rhodium	
29																					46	Pd 106.32 Palladium	
30																					47	Ag 107.87 Silver	
31																					48	Cd 112.41 Cadmium	
32																					49	In 114.82 Indium	
33																					50	Sn 118.71 Tin	
34																					51	Sb 121.76 Antimony	
35																					52	Te 127.60 Tellurium	
36																					53	I 126.91 Iodine	
37																					54	Xe 131.29 Xenon	
38																					55	Cs 132.91 Cesium	
39																					56	Ba 137.33 Barium	
40																					57 to 70		
41																					71	Lu 174.97 Lutetium	
42																					72	Hf 178.49 Hafnium	
43																					73	Ta 180.95 Tantalum	
44																					74	W 183.84 Tungsten	
45																					75	Re 186.21 Rhenium	
46																					76	Os 190.23 Osmium	
47																					77	Ir 192.22 Iridium	
48																					78	Pt 195.08 Platinum	
49																					79	Au 196.97 Gold	
50																					80	Hg 200.59 Mercury	
51																					81	Tl 204.38 Thallium	
52																					82	Pb 207.2 Lead	
53																					83	Bi 208.98 Bismuth	
54																					84	Po 209 Polonium	
55																					85	At 210 Astatine	
56																					86	Rn 222 Radon	
57 to 70																					71 to 102		
71																					103	La 138.91 Lanthanum	
72																					104	Rf 181 Rutherfordium	
73																					105	Db 185 Dubnium	
74																					106	Sg 271 Seaborgium	
75																					107	Bh 264 Bohrium	
76																					108	Hs 277 Hassium	
77																					109	Mt 268 Meitnerium	
78																					110	Ds 271 Darmstadtium	
79																					111	Rg 272 Roentgenium	
80																					112	Cn 285 Copernicium	
81																					113	Nh 284 Nihonium	
82																					114	Fl 289 Flerovium	
83																					115	Mc 288 Moscovium	
84																					116	Lv 293 Livermorium	
85																					117	Ts 294 Tennessine	
86																					118	Og 294 Oganesson	

- Group names*
- Hydrogen & the alkali metals
 - Alkaline earth metals
 - Coinage metals
 - Volatile metals
 - Isotopes
 - Chalcogens
 - Postcogens
 - Chalcogens
 - Halogens
 - Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
(Alred-Pauling)

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Actinides**
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, selenates, and tellurates.

Metals

- Metalloids**
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).
- Group 1**
The alkali metals (Li, Na, K, Rb, Cs, Fr) and the alkaline earth metals (Be, Mg, Ca, Sr, Ba, Ra) are the most reactive metals in the periodic table.

Nonmetals

- Chalcogens**
O, S, Se, and Te are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive
Oganesson is likely to be either a metalloid or a post-transition metal

Themes On The Periodic Table Answers Key

AW Rasmussen



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