

$c$  = specific heat ( $J/g \cdot ^\circ C$ )       $q$  = heat energy ( $J$ )  
 $T$  = temperature ( $^\circ C$ )       $m$  = mass ( $g$ )       $\Delta$  = change in

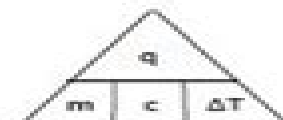
$$q = m \cdot c \cdot \Delta T$$

Name \_\_\_\_\_ Period \_\_\_\_\_

# Thermal Energy & Heat Unit Study Guide

**Part 1.** For each of the following scenarios:

- Identify as vaporization, condensation, melting, or freezing
- Write the change in states [example: gas  $\rightarrow$  liquid]
- Write whether the process absorbs or releases thermal energy



1. High altitude rain turns to hail.

a) \_\_\_\_\_ c) \_\_\_\_\_

b) \_\_\_\_\_

2. Towering cumulonimbus clouds form in the sky.

a) \_\_\_\_\_ c) \_\_\_\_\_

b) \_\_\_\_\_

3. A shallow puddle dries up over a couple day span.

a) \_\_\_\_\_ c) \_\_\_\_\_

b) \_\_\_\_\_

**Part II. Specific Heat Calculations** Refer to the equation and symbol legend above. Hint – use the triangle method. Show your work.

- A 110 g sample ( $m$ ) of tin was heated from  $23.2^\circ C$  to  $39.8^\circ C$ . What was the change in thermal energy? Solve for the heat energy ( $q$ ). Tin has a specific heat ( $c$ ) value of  $0.2274 J/g \cdot ^\circ C$ .

$c =$  \_\_\_\_\_

$m =$  \_\_\_\_\_

$\Delta T =$  \_\_\_\_\_

- A scientist heated bundled strands of copper wire from  $21.0^\circ C$  to  $82.0^\circ C$ . Copper has a specific heat ( $c$ ) capacity of  $0.3845 J/g \cdot ^\circ C$ . If the change in heat energy ( $q$ ) was 192 J, what was the mass ( $m$ ) of the bundled strands of copper wire?

$c =$  \_\_\_\_\_

$\Delta T =$  \_\_\_\_\_

$q =$  \_\_\_\_\_

# Thermal Energy And Heat Study Guide

**Patricia L. Samuel**



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