# Thermochemistry - Energy Flow and Chemical Changes

 Determine the amount of heat (in kJ) given off when 1.26x10<sup>4</sup> g of NO<sub>2</sub> are produced according to the equation:

$$2NO(g) + O_2(g) \rightarrow 2NO_2(g)$$
  $\Delta H = -114.6 \text{ kJ/mol}$ 

2. Consider the reaction:

$$H_2(g) + Cl_2(g) > 2HCl(g)$$
  $\Delta H = -184.6 \text{ kJ/mol}$   
If 3 moles of  $H_2$  react with 3 moles of  $Cl_2$  to form  $HCl_c$  calculate the work done (in joules) against a pressure of 1.0 atm at 25 °C. What is  $\Delta E$  for this reaction? Assume the reaction goes to completion.

 A quantity of 200. mL of 0.862 M HCl is mixed with 200. mL of 0.431 M Ba(OH)<sub>2</sub> in a constantpressure calorimeter of negligible heat capacity. The initial temperature of the HCl and Ba(OH)<sub>2</sub> solutions is the same at 20.48 °C. For the process:

$$H^*(aq) + OH^*(aq) \rightarrow H_2O(l)$$
  $\Delta H = -56.2 \text{ kJ/mol.}$ 

What is the final temperature of the mixed solution? (Assume that the combined solution has a mass of 400g and a specific heat capacity of 4.184 J/°C·g)

A. Eroom the following data

# **Thermochemistry Heat And Chemical Changes Answers**

**Christian G. Meyer** 

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