

Thermochemistry Practice Problems

- 10) How much work in joules is done when a piston expands from a volume of 13.27 liters to 76.55 liters against a pressure of 14.89 atm?

10) _____

$$-15450 \text{ J}$$

- 11) How much work in joules is done on the system when a 1.15 atm external pressure causes a piston to decrease in volume from 6.55 liters to 3.16 liters?

11) _____

$$395 \text{ J}$$

- 12) In a refrigeration system, the refrigerant gas absorbs 21.39 kJ of energy while expanding against a 0.278 atmosphere pressure from a volume of 0.0423 liters to a volume of 1.876 liters. What is the energy change of the gas?

12) _____

$$21.338 \text{ kJ}$$

- 13) What is the energy change of the system if a chemical reaction transfers 32.146 kJ of heat to the surroundings while it causes the expansion of a 1.465 liter vessel to 3.687 liters against a pressure of 3.64 atmospheres?

13) _____

$$\Delta E = -32.965 \text{ kJ}$$

- 14) Lead, water, sulfur, and arsenic have specific heats of 0.128, 4.18, 0.706, and 0.329 J/g°C, respectively. Which of these would require the smallest amount of heat to increase its temperature by 10 °C (assume all samples have the same mass)?

14) _____

LEAD $\Rightarrow$ SMALLEST SPECIFIC HEAT

- 15) When power was turned off to a 30.0 gal. water heater, the temperature of the water dropped from 75.0 °C to 22.5 °C. How much heat was lost to the surroundings? (1 gal = 3.785 L)

15) _____

$$2.49 \times 10^4 \text{ kJ}$$

Thermochemistry Problem Set One Specific Heat

DP Hallahan



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Table of Contents Thermochemistry Problem Set One Specific Heat

1. Understanding the eBook Thermochemistry Problem Set One Specific Heat
 - The Rise of Digital Reading Thermochemistry Problem Set One Specific Heat
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Problem Set One Specific Heat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Problem Set One Specific Heat
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Problem Set One Specific Heat
 - Personalized Recommendations

- Thermochemistry Problem Set One Specific Heat User Reviews and Ratings
- Thermochemistry Problem Set One Specific Heat and Bestseller Lists
- 5. Accessing Thermochemistry Problem Set One Specific Heat Free and Paid eBooks
 - Thermochemistry Problem Set One Specific Heat Public Domain eBooks
 - Thermochemistry Problem Set One Specific Heat eBook Subscription Services
 - Thermochemistry Problem Set One Specific Heat Budget-Friendly Options
- 6. Navigating Thermochemistry Problem Set One Specific Heat eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Problem Set One Specific Heat Compatibility with Devices
 - Thermochemistry Problem Set One Specific Heat Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Problem Set One Specific Heat
 - Highlighting and Note-Taking Thermochemistry Problem Set One Specific Heat
 - Interactive Elements Thermochemistry Problem Set One Specific Heat
- 8. Staying Engaged with Thermochemistry Problem Set One Specific Heat
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Problem Set One Specific Heat
- 9. Balancing eBooks and Physical Books Thermochemistry Problem Set One Specific Heat
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Problem Set One Specific Heat
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Problem Set One Specific Heat
 - Setting Reading Goals Thermochemistry Problem Set One Specific Heat
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Problem Set One Specific Heat
 - Fact-Checking eBook Content of Thermochemistry Problem Set One Specific Heat

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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