

3.3 Thermodynamics: Study Guide

Use this [outline](#) to review the key vocabulary for this unit.

1. Match the vocabulary words on the left with the appropriate definitions and units on the right.

a. Thermal Energy	<u>A</u> Kinetic energy in transit from one object to another due to temperature difference
b. Temperature	<u>B</u> Degree Celsius
	<u>C</u> Joules
	<u>D</u> The average kinetic energy of particles in an object.

2. In the table below, write in the appropriate formulas.

Energy Transfer	Rate of Energy Transfer	Thermal Conductivity	Coefficient of Heat Conductivity
$Q = mc \cdot \Delta T$	$P = \frac{c \Delta T}{L}$ or $P = \frac{Q}{t}$ <u>Thermal conductance</u> $P = \sigma A c (\Delta T)^2$	$K = \frac{PL}{A \Delta T}$	$U = \frac{P}{A \cdot \Delta T}$

3. For each example, write the correct number for the law of thermodynamics it describes (0, 1, 2).
- 2 The law of entropy
 - 1 Thermal energy can only be increased in a system by adding heat from an external source or performing work within the system
 - 0 If two systems are separately found to be in thermal equilibrium with a third, then the two systems are in thermal equilibrium with each other.
 - 2 Thermal energy flows only from warmer to cooler
 - 1 The law of conservation of energy applied to thermal systems
4. How much thermal energy would it take to raise the temperature of 1.1 (1 kg) of water by 10°C? The specific heat capacity of water is $4184 \frac{J}{kg \cdot ^\circ C}$

$$Q = mc \Delta T$$

$$= 1.1 kg \cdot 4184 \frac{J}{kg \cdot ^\circ C} \cdot 10^\circ C$$

$$Q = 41,840 J$$

$$Q = \underline{41,840 J}$$

5. A materials scientist is trying to determine the specific heat capacity of an experimental material. She applies 30 kJ of thermal energy to a 5 kg sample of the material and records a 24°C temperature increase in the sample. What is the specific heat capacity of the material?

$$Q = mc \Delta T$$

$$c = \frac{Q}{m \Delta T}$$

$$c = 2500 \frac{J}{kg \cdot ^\circ C}$$

$$c = \underline{2500 \frac{J}{kg \cdot ^\circ C}}$$

$$= \frac{30000 J}{5 kg \cdot 24^\circ C}$$

Thermodynamics Study Guide Answers

Ian Pickup



Thermodynamics Study Guide Answers:

The book delves into Thermodynamics Study Guide Answers. Thermodynamics Study Guide Answers is a crucial topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Thermodynamics Study Guide Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Thermodynamics Study Guide Answers
 - Chapter 2: Essential Elements of Thermodynamics Study Guide Answers
 - Chapter 3: Thermodynamics Study Guide Answers in Everyday Life
 - Chapter 4: Thermodynamics Study Guide Answers in Specific Contexts
 - Chapter 5: Conclusion
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