

### Unit 3 Review: Thermochemistry

**A. Definition Questions:** you should be able to use or apply the following terms. Be able to write complete definitions for the terms in **bold**.

energy  
potential energy  
kinetic energy  
**enthalpy**  
thermal kinetic energy

temperature  
**heat**  
**specific heat capacity**  
endothermic  
exothermic

standard state  
allotropes  
**standard heat of formation**  
**bond energy**

- Identify the **INCORRECT** statement below:
  - potential energy is the energy an object has because of its position relative to and attraction toward another object.
  - energy is the ability to do work.
  - in an exothermic reaction, the enthalpy of the products is higher than the enthalpy of the reactants.
  - in an ordinary chemical reaction, energy is neither created nor destroyed.
- A particular chemical reaction has  $\Delta H^\circ = +250 \text{ kJ/mol}$ . Which of the following statements is/are true concerning this reaction?
  - Heat moves from the system to the surroundings.
  - The reaction is endothermic.
  - The enthalpy of the products is lower than the enthalpy of the reactants.
  - I and II only
  - II only
  - II and III only
  - I, II and III are true
- Which of the following molecules has the highest enthalpy?
  - propane
  - cyclopropane
  - pentane
  - cyclooctane
- Which statement is **INCORRECT** about endothermic reactions?
  - the system absorbs energy from its surroundings
  - the enthalpy of products is lower than the enthalpy of the reactants
  - the thermal kinetic energy of the surroundings will decrease
  - the enthalpy change will have a positive value
- All of the following statements are true **EXCEPT**:
  - $\Delta H$  for a reaction is equal in magnitude but opposite in sign to  $\Delta H$  for the reverse reaction
  - enthalpy is a state function
  - the overall change in enthalpy for a reaction depends on the number of steps in a reaction
  - chemical potential energy contributes to the total enthalpy of a substance
- What is the definition of the temperature of a substance?
  - the total heat content of all the particles in a substance
  - the speed of the fastest particles in the substance
  - the speed of the slowest particles in the substance
  - the average kinetic energy of the particles in a system

# Unit 13 Test Review Thermochemistry

**Huangqi Zhang**



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