

Sample Test for Unit VII

Potentially useful equations:

$$Q = mc\Delta T \quad Q = mL \quad Q = mH_c \quad \varepsilon_c = 1 - \frac{T_c}{T_H} \quad \varepsilon = \frac{W}{Q_{in}}$$

Specific heat for water = $4186 \text{ J kg}^{-1}\text{C}^{-1}$

1. a) Define the kilocalorie.
b) How much heat energy in kilocalories is required to change 2.3 kg of water from 38°C to 45°C ?
2. a) Define Heat of Combustion.
b) A fuel-oil burning boiler is used to heat water from 20°C to 60°C . The Heat of Combustion for fuel-oil is $4.19 \times 10^7 \text{ J/kg}$. The boiler is 40% efficient. How much fuel-oil is required to heat 5000 kg of water?
3. Fifty grams of glycerine at 35.0°C is mixed with 1.00 kg of water at 20.0°C . Find the specific heat of glycerine if the equilibrium temperature is measured to be 20.42°C ; assume no heat losses or gains.
4. a) Sketch a typical phase diagram with two phase changes, labeling the axes and transitions. (Also think about what determines the slopes of the T - Q lines: How are they related to the specific heat capacity?)
b) Find the total heat energy required to change 2.0 kg of ice at -10°C into liquid water at 20°C . (Specific heat for ice = $2100 \text{ J kg}^{-1}\text{C}^{-1}$, Latent Heat of Fusion = $3.34 \times 10^5 \text{ J/kg}$, Specific heat for water = $4186 \text{ J kg}^{-1}\text{C}^{-1}$, Latent Heat of Vaporization = $22.6 \times 10^5 \text{ J/kg}$).
5. Define and provide an example for each of the three methods of heat transfer.
6. a) State the First Law of Thermodynamics.
b) State the Second Law of Thermodynamics.
7. a) Draw and label a schematic diagram of a heat engine.
b) Determine the Carnot efficiency of a steam turbine engine which operates between a high temperature of 550°C and a low temperature of 90°C .

Short Answers:

1. a) The kcal is the heat energy needed to raise 1 kg of water by 1°C (from 14.5°C to 15.5°C).
b) $Q = mc\Delta T$, where $c_{\text{water}} = 1 \text{ kcal kg}^{-1}\text{C}^{-1}$ ($= 4186 \text{ J kg}^{-1}\text{C}^{-1}$), so that
 $Q = (2.3 \text{ kg})(1 \text{ kcal kg}^{-1}\text{C}^{-1})(45^\circ\text{C} - 38^\circ\text{C}) = 16.1 \text{ kcal}$.

Unit Vii Review Answers Physics

Didier Musso



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Table of Contents Unit Vii Review Answers Physics

1. Understanding the eBook Unit Vii Review Answers Physics
 - The Rise of Digital Reading Unit Vii Review Answers Physics
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit Vii Review Answers Physics
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit Vii Review Answers Physics
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit Vii Review Answers Physics
 - Personalized Recommendations
 - Unit Vii Review Answers Physics User Reviews and Ratings
 - Unit Vii Review Answers Physics and Bestseller Lists

5. Accessing Unit Vii Review Answers Physics Free and Paid eBooks
 - Unit Vii Review Answers Physics Public Domain eBooks
 - Unit Vii Review Answers Physics eBook Subscription Services
 - Unit Vii Review Answers Physics Budget-Friendly Options
6. Navigating Unit Vii Review Answers Physics eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit Vii Review Answers Physics Compatibility with Devices
 - Unit Vii Review Answers Physics Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit Vii Review Answers Physics
 - Highlighting and Note-Taking Unit Vii Review Answers Physics
 - Interactive Elements Unit Vii Review Answers Physics
8. Staying Engaged with Unit Vii Review Answers Physics
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit Vii Review Answers Physics
9. Balancing eBooks and Physical Books Unit Vii Review Answers Physics
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit Vii Review Answers Physics
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit Vii Review Answers Physics
 - Setting Reading Goals Unit Vii Review Answers Physics
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit Vii Review Answers Physics
 - Fact-Checking eBook Content of Unit Vii Review Answers Physics
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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