Date Pd _____Pd

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

Write the <u>balanced chemical equation</u>.

Equation: 2AL +6HCL →2ALCL3 + 3H2

b. Identify what is given (with units) and what you want to find (with units).

Use coefficients from balanced equation to determine mole ratio.

d. Show set up (organize it!).

 Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation:	$_2$ $_1$ $_2$ $_3$ $_4$ $_5$ $_4$ $_5$ $_5$ $_5$ $_5$ $_5$ $_6$ $_7$ $_7$ $_7$ $_7$ $_7$ $_7$ $_7$ $_7$				
Before	_2_	_ex	_0	_o	3.0 mol O2
Change	2	3	_2	_2	
After	o	ex_	_2	2_	

Propane, C3Hs, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Before 0.5O O 1.50mol HCL 0.00 Change -0.5-1.5+0.5 ± 0.5 After 0 68000 0.05 0.75

3. Ammonia, NH3, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: 3H2 + N2 → 2NH3

Before ex 1.25 0 2.50 mole

Change -3.75 -1.25 +25

After ex 0 2.5

Unit 8 I Mole Relationships Answers

Jicheng Xie

Unit 8 I Mole Relationships Answers:

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