

Unit 7 Chemistry - Nail Lab

Purpose *After you have completed the lab, think about what was the overall point for performing the lab. What were we trying to accomplish?*

Procedure

Day 1

1. Mass a clean & labeled _____ mL beaker.
2. Put between _____ g of copper (II) chloride in the beaker. Record the mass of the beaker and the copper (II) chloride.
3. Add _____ mL of distilled water to the beaker. Stir to dissolve the solid.
4. Obtain the mass of _____ iron nails together to ± 0.01 g.
5. Place the nails in the copper (II) chloride solution. Observe the reaction; record your observations. Place the labeled beaker in the place designated by your teacher.

Day 2

6. Remove the nails. Rinse or scrape the _____ (copper metal) from the nails into your labeled beaker. Place the nails in a separate labeled small beaker. Note the appearance of the nails in your observations. Place this beaker in the drying oven.
7. _____ the solution from the 250 mL beaker. Rinse the precipitate with about _____ mL of distilled water. Try to lose as little of the solid copper as you can when you decant. After a 2nd rinse with distilled water, rinse the copper with _____. Rinse one last time with distilled water. Then place the labeled beaker in the drying oven.
8. Mass the dry nails, and then discard the nails.

Day 3

9. Mass the beaker + dry copper. Discard the copper in the place designated by your teacher. Wash your beaker and let dry.

Data:

Mass 250 mL beaker	
Mass 250 mL beaker + copper (II) chloride	
Mass nails before reaction	
Mass nails after reaction	

Unit Nail Lab Chemistry Conclusion Answers

Christian G. Meyer



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