

## Sample Test for Unit VII

Potentially useful equations:

$$Q = mc\Delta T \quad Q = mL \quad Q = mH_c \quad \varepsilon_c = 1 - \frac{T_C}{T_H} \quad \varepsilon = \frac{W}{Q_{in}}$$

Specific heat for water =  $4186 \text{ J kg}^{-1}\text{C}^{-1}$

1. a) Define the kilocalorie.  
b) How much heat energy in kilocalories is required to change 2.3 kg of water from  $38^\circ\text{C}$  to  $45^\circ\text{C}$ ?
2. a) Define Heat of Combustion.  
b) A fuel-oil burning boiler is used to heat water from  $20^\circ\text{C}$  to  $60^\circ\text{C}$ . The Heat of Combustion for fuel-oil is  $4.19 \times 10^7 \text{ J/kg}$ . The boiler is 40% efficient. How much fuel-oil is required to heat 5000 kg of water?
3. Fifty grams of glycerine at  $35.0^\circ\text{C}$  is mixed with 1.00 kg of water at  $20.0^\circ\text{C}$ . Find the specific heat of glycerine if the equilibrium temperature is measured to be  $20.42^\circ\text{C}$ ; assume no heat losses or gains.
4. a) Sketch a typical phase diagram with two phase changes, labeling the axes and transitions. (Also think about what determines the slopes of the  $T$ - $Q$  lines: How are they related to the specific heat capacity?)  
b) Find the total heat energy required to change 2.0 kg of ice at  $-10^\circ\text{C}$  into liquid water at  $20^\circ\text{C}$ . (Specific heat for ice =  $2100 \text{ J kg}^{-1}\text{C}^{-1}$ , Latent Heat of Fusion =  $3.34 \times 10^5 \text{ J/kg}$ , Specific heat for water =  $4186 \text{ J kg}^{-1}\text{C}^{-1}$ , Latent Heat of Vaporization =  $22.6 \times 10^5 \text{ J/kg}$ ).
5. Define and provide an example for each of the three methods of heat transfer.
6. a) State the First Law of Thermodynamics.  
b) State the Second Law of Thermodynamics.
7. a) Draw and label a schematic diagram of a heat engine.  
b) Determine the Carnot efficiency of a steam turbine engine which operates between a high temperature of  $550^\circ\text{C}$  and a low temperature of  $90^\circ\text{C}$ .

### Short Answers:

1. a) The kcal is the heat energy needed to raise 1 kg of water by  $1^\circ\text{C}$  (from  $14.5^\circ\text{C}$  to  $15.5^\circ\text{C}$ ).  
b)  $Q = mc\Delta T$ , where  $c_{\text{water}} = 1 \text{ kcal kg}^{-1}\text{C}^{-1}$  ( $= 4186 \text{ J kg}^{-1}\text{C}^{-1}$ ), so that  
 $Q = (2.3 \text{ kg})(1 \text{ kcal kg}^{-1}\text{C}^{-1})(45^\circ\text{C} - 38^\circ\text{C}) = 16.1 \text{ kcal}$ .

# Unit Vii 6 Answers Physics

**Leonard Vincent Koos**



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## **Table of Contents Unit Vii 6 Answers Physics**

1. Understanding the eBook Unit Vii 6 Answers Physics
  - The Rise of Digital Reading Unit Vii 6 Answers Physics
  - Advantages of eBooks Over Traditional Books
2. Identifying Unit Vii 6 Answers Physics
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Unit Vii 6 Answers Physics
  - User-Friendly Interface
4. Exploring eBook Recommendations from Unit Vii 6 Answers Physics
  - Personalized Recommendations
  - Unit Vii 6 Answers Physics User Reviews and Ratings

- Unit Vii 6 Answers Physics and Bestseller Lists
- 5. Accessing Unit Vii 6 Answers Physics Free and Paid eBooks
  - Unit Vii 6 Answers Physics Public Domain eBooks
  - Unit Vii 6 Answers Physics eBook Subscription Services
  - Unit Vii 6 Answers Physics Budget-Friendly Options
- 6. Navigating Unit Vii 6 Answers Physics eBook Formats
  - ePub, PDF, MOBI, and More
  - Unit Vii 6 Answers Physics Compatibility with Devices
  - Unit Vii 6 Answers Physics Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Unit Vii 6 Answers Physics
  - Highlighting and Note-Taking Unit Vii 6 Answers Physics
  - Interactive Elements Unit Vii 6 Answers Physics
- 8. Staying Engaged with Unit Vii 6 Answers Physics
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Unit Vii 6 Answers Physics
- 9. Balancing eBooks and Physical Books Unit Vii 6 Answers Physics
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Unit Vii 6 Answers Physics
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Unit Vii 6 Answers Physics
  - Setting Reading Goals Unit Vii 6 Answers Physics
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit Vii 6 Answers Physics
  - Fact-Checking eBook Content of Unit Vii 6 Answers Physics
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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