

Worksheet 1 – Scientific Notation and Significant Figures

1. Convert each of the following into scientific notation.

- | | | |
|---------------------------|-------------------------|--------------------------|
| a) 3427 | b) 0.00456 | c) 123,453 |
| d) 172 | e) 0.000984 | f) 0.502 |
| g) 3100.0×10^2 | h) 0.0114×10^4 | i) 107.2 |
| j) 0.0000455 | k) 2205.2 | l) 30.0×10^{-2} |
| m) 0.982×10^{-3} | n) 0.0473 | o) 650.502 |
| p) 3.03×10^{-1} | q) 20.4×10^5 | r) 1.29 |
| s) 0.00565 | t) 1362205.2 | u) 450.0×10^3 |
| v) 1000×10^{-3} | | |

2. Determine the number of significant figures in each of the following:

- | | | |
|---------------------------|-------------------------|--------------------------|
| a) 3427 | b) 0.00456 | c) 123,453 |
| d) 172 | e) 0.000984 | f) 0.502 |
| g) 3100.0×10^2 | h) 0.0114×10^4 | i) 107.2 |
| j) 0.0000455 | k) 2205.2 | l) 30.0×10^{-2} |
| m) 0.982×10^{-3} | n) 0.0473 | o) 650.502 |
| p) 3.03×10^{-1} | q) 20.4×10^5 | r) 1.29 |
| s) 0.00565 | t) 1362205.2 | u) 450.0×10^3 |
| v) 1000×10^{-3} | w) $546,000 \pm 10$ | x) $546,000 \pm 1000$ |

3. Convert each into decimal form.

- | | | | |
|-----------------------|---------------------------|------------------------|---------------------|
| 1.56×10^4 | 0.56×10^{-2} | 3.69×10^{-2} | 736.9×10^5 |
| 0.00259×10^5 | 0.000459×10^{-1} | 13.69×10^{-2} | 6.9×10^4 |
| 0.00259×10^3 | 0.0209×10^{-3} | | |

4. Calculate the following. Give the answer in correct scientific notation.

- | | | | |
|--|-----------------------------------|---|--|
| a) 4.53×10^5
+ 2.2×10^6 | b) 1913.0
- 4.6×10^3 | c) 2.34×10^{-4}
+ 1.92×10^{-23} | d) 2.130×10^3
- 6.6×10^2 |
| e) 9.10×10^3
+ 2.2×10^6 | f) 1113.0
- 14.6×10^2 | g) 6.18×10^{-45}
+ 4.72×10^{-44} | h) 4.25×10^{-3}
- 1.6×10^{-2} |

5. Calculate the following. Give the answer in correct scientific notation.

- a) $3.95 \times 10^2 / 1.5 \times 10^6$ b) $(3.5 \times 10^2)(6.45 \times 10^{10})$ c) $4.44 \times 10^7 / 2.25 \times 10^5$ d) $(4.50 \times 10^{-12})(3.67 \times 10^{-12})$
- e) $1.05 \times 10^{-26} / 4.2 \times 10^{-26}$ f) $(2.5 \times 10^9)(6.45 \times 10^4)$ g) $6.022 \times 10^{23} / 3.011 \times 10^{-26}$ h) $(6.68 \times 10^5)(3.45 \times 10^{-10})$

Unit 08 Practice Problems Chemistry Key Significant Figures

Kenneth W. Raymond



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Chemistry: 1001 Practice Problems For Dummies (+ Free Online Practice) Heather Hattori, Richard H. Langley, 2022-05-10 Practice your way to a better grade in your Chemistry class Chemistry 1001 Practice Problems For Dummies gives you 1 001 opportunities to practice solving problems on all the topics covered in your chemistry class in the book and online Get extra practice with tricky subjects solidify what you ve already learned and get in depth walk throughs for every problem with this useful book These practice problems and detailed answer explanations will catalyze the reactions in your brain no matter what your skill level Thanks to Dummies you have a resource to help you put key concepts into practice Work through multiple choice practice problems on all Chemistry topics covered in class Step through detailed solutions to build your understanding Access practice questions online to study anywhere any time Improve your grade and up your study game with practice practice practice The material presented in Chemistry 1001 Practice Problems For Dummies is an excellent resource for students as well as parents and tutors looking to help supplement classroom instruction Chemistry 1001 Practice Problems For Dummies 9781119883531 was previously published as 1 001 Chemistry Practice Problems For Dummies 9781118549322 While this version features a new Dummies cover and design the content is the same as the prior release and should not be considered a new or updated product **CliffsNotes Chemistry Practice Pack** Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie s classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie s many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value **Basic Concepts of Chemistry** Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from

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in Chemistry, Second Edition Daley,1988-02-19 **General, Organic, and Biological Chemistry** Kenneth W.

Raymond,2013-01-04 General Organic and Biological Chemistry 4th Edition has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry An integrated approach is employed in which related general chemistry organic chemistry and biochemistry topics are presented in adjacent chapters This approach helps students see the strong connections that exist between these three branches of chemistry and allows instructors to discuss these interrelationships while the material is still fresh in students minds **Foundations of College Chemistry** Morris Hein,Susan Arena,2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry

Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis Medical Laboratory

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Concepts And Problems In Physical Chemistry P.S. Raghavan,1997 Contents Introduction Atoms Molecules and Formulas Chemical Equations and Stoichiometry Aqueous Reactions and Solution Stoichiometry Gases Intermolecular Forces Liquids and Solids Atoms Structure and the Periodic Table Chemical Bonding Chemical Thermodynamics Solutions Chemical Kinetics Chemical Equilibrium Acids and Bases Ionic Equilibria I Ionic Equilibria II Redox Reactions Electrochemistry Nuclear Chemistry **eBook: General, Organic and Biological Chemistry 2e** SMITH,2012-02-16 eBook General Organic and Biological Chemistry 2e **Analytical Chemistry, International Adaptation** Gary D. Christian,Purnendu K.

Dasgupta, Kevin A. Schug, 2020 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry Comprehensive Chemistry XI Dr. B. Kapila, S. K. Khanna, 2010-11 Comprehensive chemistry according to the new syllabus prescribed by Central Board of Secondary Education CBSE Investigating Chemistry Matthew Johll, 2008-12-22 In its new second edition Investigating Chemistry A Forensic Science Perspective remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package **Ebook: Chemistry: The Molecular Nature of Matter and Change** Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change *Chemistry: Core Concepts, 3rd Edition* Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package *Chemistry* Carson-Dellosa Publishing, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards **Foundations of College Chemistry, Alternate** Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped

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Fundamentals of Chemistry: A Modern Introduction (1966) Frank Brescia, 2012-12-02 Fundamentals of Chemistry A Modern Introduction focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro s Law and Graham s Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to study chemistry

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web nov 25 2017 however the german version of the short form of the social phobia inventory mini spin has not been evaluated yet therefore our aim was to determine reliability validity and population based norms of the german mini spin

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