

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide CuO
2. mercury (I) oxide Hg₂O
3. gold (III) chloride AuCl₃
4. thallium (III) bromide TlBr₃
5. bismuth (V) oxide Bi₂O₅
6. terbium (III) oxide Tb₂O₃
7. uranium (VI) oxide UO₃
8. protactinium (IV) bromide PaBr₄
9. cerium (III) oxide Ce₂O₃
10. arsenic (V) sulphide As₂S₅

11. manganese (III) oxide Mn₂O₃
12. vanadium (II) bromide VBr₂
13. niobium(V) oxide Nb₂O₅
14. titanium (III) oxide Ti₂O₃
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co₃P₂
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr₂
20. copper (II) bromide CuBr₂

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl₄ Tin (IV) Chloride
2. BiBr₅ Bismuth (V) Bromide
3. PoO₂ Polonium (IV) Oxide
4. PbI₂ Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au₂O₃ Gold (III) Oxide
8. FeCl₂ Iron (II) Chloride

9. PdF₄ Palladium (IV) Fluoride
10. Os₂O₃ Osmium (III) Oxide
11. MoBr₂ Molybdenum (II) Bromide
12. VCl₅ Vanadium (V) Chloride
13. Mn₂O₃ Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np₂O₃ Neptunium (III) Oxide
16. V₂O₅ Vanadium (V) Oxide

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Table of Contents Unit 10 Bonding And Formula Writing Answers

1. Understanding the eBook Unit 10 Bonding And Formula Writing Answers
 - The Rise of Digital Reading Unit 10 Bonding And Formula Writing Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 10 Bonding And Formula Writing Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 10 Bonding And Formula Writing Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 10 Bonding And Formula Writing Answers
 - Personalized Recommendations
 - Unit 10 Bonding And Formula Writing Answers User Reviews and Ratings
 - Unit 10 Bonding And Formula Writing Answers and Bestseller Lists

5. Accessing Unit 10 Bonding And Formula Writing Answers Free and Paid eBooks
 - Unit 10 Bonding And Formula Writing Answers Public Domain eBooks
 - Unit 10 Bonding And Formula Writing Answers eBook Subscription Services
 - Unit 10 Bonding And Formula Writing Answers Budget-Friendly Options
6. Navigating Unit 10 Bonding And Formula Writing Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 10 Bonding And Formula Writing Answers Compatibility with Devices
 - Unit 10 Bonding And Formula Writing Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 10 Bonding And Formula Writing Answers
 - Highlighting and Note-Taking Unit 10 Bonding And Formula Writing Answers
 - Interactive Elements Unit 10 Bonding And Formula Writing Answers
8. Staying Engaged with Unit 10 Bonding And Formula Writing Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 10 Bonding And Formula Writing Answers
9. Balancing eBooks and Physical Books Unit 10 Bonding And Formula Writing Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 10 Bonding And Formula Writing Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit 10 Bonding And Formula Writing Answers
 - Setting Reading Goals Unit 10 Bonding And Formula Writing Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 10 Bonding And Formula Writing Answers
 - Fact-Checking eBook Content of Unit 10 Bonding And Formula Writing Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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