

6.023×10^{23}
particles

in terms of
particles

in terms of
volume

22.4 L of a
gas at N.T.P

1 Mole

1 mole of oxygen atoms
= 6.023×10^{23} atoms

1 mole of oxygen molecules
= 6.023×10^{23} molecules

in terms of
mass

1 gram atom of
elements

1 gram molecule
of substance

1 gram formula
weight of
substance

12g
C

23g
Na

18g
H₂O

44g
CO₂

58.5 g
NaCl

100g
CaCO₃

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L Towner

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