

UNIT VII Review Stoichiometry

Stoichiometry Calculations involving MASS, MOLES, VOLUME, and MOLECULES

1. Consider the reaction $4\text{NH}_3 + 5\text{O}_2 \rightarrow 6\text{H}_2\text{O} + 4\text{NO}$
 - a. What mass of NO is produced when 5.00 mol of NH_3 are reacted with excess O_2 ? 150g NO
 - b. What volume of O_2 at STP is required to produce 90.0L of NO? 113L O_2
 - c. How many molecules of H_2O are formed when 100.0g of NH_3 react? 5.31×10^{24} molecules H_2O
2. Consider the reaction $\text{SiCl}_4 + 2\text{H}_2 \rightarrow \text{Si} + 4\text{HCl}$
 - a. How many moles of SiCl_4 are required to produce 4.25×10^{12} molecules of HCl. 1.76×10^{-12} mol SiCl_4
 - b. If 25.0 mL of H_2 (g) at STP react, calculate the mass of Si produced. 0.0157g Si
 - c. If 0.892 kg of SiCl_4 react, calculate the number of molecules of HCl produced. 21.0 mol HCl
3. Silver reacts with nitric acid, HNO_3 to form nitrogen monoxide, NO, silver nitrate, AgNO_3 and water.

O M I T

 - a. How many grams of nitric acid are required to react with 5.00g of silver?
 - b. How many grams of silver nitrate will be formed?
 - c. How many moles of nitrogen monoxide will be produced?

Stoichiometry using MOLAR CONCENTRATION

4. Consider the following reaction $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
 - a. What volume of HCl having a concentration of 0.0010M is neutralized by a 0.850 g portion of CaCO_3 ? 17.0L HCl
 - b. What volume of CO_2 at STP is produced if 1.25L of 0.0055 M HCl reacts with an excess of CaCO_3 ? 0.077L CO_2
 - c. How many moles of 0.200 M HCl is reacted to produce 3.00g of CaCl_2 ? 0.024 mol H_2
5. If 17.5 g of zinc are reacted with phosphoric acid, H_3PO_4 , then zinc phosphate, $\text{Zn}_3(\text{PO}_4)_2$ and hydrogen are produced.
 - a. How many moles of phosphoric acid are required? 0.178 mol H_3PO_4
 - b. If the phosphoric acid is a 3.00M solution, how many litres of the solution are needed? 0.0593L
 - c. What mass of zinc phosphate will be produced? 84.4g $\text{Zn}_3(\text{PO}_4)_2$

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