

Now _____ Day _____ One _____

CHAPTER 9 **STUDY GUIDE FOR CONTENT MASTERY**

Covalent Bonding

Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	p bond
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When sharing of electron occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5) p bond**.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is what form the elements such as hydrogen, nitrogen, and oxygen normally exist?
 - a single atom
 - (b) in molecules containing two atoms**
 - in molecules containing three atoms
 - in molecules containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - (d) four**
- Bond length is the distance between
 - two molecules of the same substance
 - (b) the electron in two attached atoms**
 - the electron in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - (a) the shorter the bond, the stronger the bond**
 - the shorter the bond, the weaker the bond
 - the shorter the bond, the lower the electron in it
 - the shorter the bond, the lower the bond dissociation energy

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Chemistry: Matter and Change • Chapter 9

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Now _____ Day _____ One _____

CHAPTER 9 **STUDY GUIDE FOR CONTENT MASTERY**

Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

- Binary molecular compounds are generally composed of a metal and a nonmetal. **false**
- The second element in the formula of a binary compound is named using the suffix -ite. **false**
- The prefix zero indicates three atoms. **false**
- The prefix four indicates six atoms. **true**
- In naming the first element in a formula, the prefix zero is not used. **true**
- For binary acids, the hydrogen part of the compound is named using the prefix hydro. **true**
- An oxyacid contains only two elements. **false**
- If the name of the anion of an oxyacid ends in -ite, the acid name contains the suffix -ous. **false**

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A

- C_2H_4
- H_2CO_3
- H_2CO
- H_2SO_4
- H_2SO_3
- $\text{H}_2\text{N}_2\text{O}_2$
- HNO_3
- HNO_2
- H_2SO_4
- H_2SO_3

Column B

- hydrobromic acid
- nitrogen tetroxide
- carbon monoxide
- nitrous acid
- ammonia
- acetic acid
- carbonic acid
- bromic acid
- carbon dioxide

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Study Guide For Content Mastery Answer Key Covalent Bonding

Daniel C. Pantaleo



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Prentice Hall Chemistry ,2000 **The Software Encyclopedia** ,1988 *School Library Journal* ,1987

Engineered Materials Abstracts ,1995-04 Chemistry Richard Post,Chad Snyder,Clifford C. Houk,2020-08-20 A

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Study Guide to Accompany Chemistry Kenneth J. Hughes,Paul B. Kelter,2002-05 Reviews essential points from the text presents problems and hints to assist the student in solving them Solutions to selected problems in the end of chapter problem sets are included By Kenneth Hughes of the University of Wisconsin Madison Study Guide [to Accompany]

General Chemistry James E. Brady,1982

Study Guide to Accompany Chemical Principles, Properties, and

Reactions Kenneth L. Henold,1984

Student Study Guide/solutions Manual to Accompany Foundations of

Inorganic, Organic, and Biological Chemistry Katherine J. Denniston,1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface

Study guide to accompany chemistry Karen Timberlakes,1983

Study Guide [to Accompany] General

Chemistry Dixie Goss,2002

Chemistry, Student Study Guide James E. Brady,Fred Senese,2008-01-28 The image on the

front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon

buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Student Study Guide to accompany Chemistry Martin Silberberg,Libby Weberg,2005-01-06 Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry Daniel C. Pantaleo,1986 Study Guide for Hein and Arena's Foundations of College Chemistry Peter Scott,Rachael Enriquez Porter,1999-04-18 *Study Guide to accompany Basic Concepts of Chemistry, 7th Edition* Leo J. Malone,2003-02-20 Work more effectively and gauge your progress along the way This Study Guide that is designed to accompany Malone s Chemistry 7th Edition includes chapter summaries new terms self tests answers to self tests and solutions to selected problems This easy to read introduction presents chemistry as a living relevant science Chemistry 7th Edition encourages critical thinking and helps readers overcome the math difficulties that often prevent them from developing a full understanding of the subject This new seventh edition builds on its core strengths of pedagogy driving the connections between ideas mathematics in context not just an appendix and an extensive problem solving emphasis with an updated design and more molecular art In addition the seventh edition expands its applications and online options One of the briefest books in the market it still provides sufficient depth for the basic concepts of chemistry

Study Guide with Solutions to Accompany Burgoyne, A Study Course in Organic Chemistry N. Sue Barnes,1979 **Study Guide and Solutions Manual to accompany Basic Concepts of Chemistry, 9th Edition** Leo J. Malone,2011-12-07 This is the Study Guide and Solutions Manual to accompany Malone s Basic Concepts of Chemistry Study Guide to Accompany Sienko and Plane[']s] Chemistry, Principles and Properties John Blair Russell,1967

The book delves into Study Guide For Content Mastery Answer Key Covalent Bonding. Study Guide For Content Mastery Answer Key Covalent Bonding is a vital topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Study Guide For Content Mastery Answer Key Covalent Bonding, encompassing both the fundamentals and more intricate discussions.

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