

Gas Laws Study Guide

Write the Law, Formula, Symbols, Units and a sentence describing the meaning of the symbols used.

1. A sample of gas is at 10.0°C and 1.00 atm.

2. The pressure is doubled and the volume is halved.

3. A sample of gas is at 10.0°C and 1.00 atm.

4. A sample of gas is at 10.0°C and 1.00 atm.

5. The volume is halved and the temperature is doubled.

6. A sample of gas is at 10.0°C and 1.00 atm.

Name _____ Date _____

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Standard Temperature = _____ °C = _____ K

Standard Pressure = _____ atm = _____ mmHg = _____ kPa = _____ torr = _____ psi

Standard Volume = _____ L = _____ mL

Write the Formula for the Gas Law:

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Combined Gas Law

$$P_{\text{Total}} = P_1 + P_2 + P_3$$

Charles' Law

$$P_1 V_1 = P_2 V_2$$

Boyle's Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Gay-Lussac's Law

$$P_1 T_1 = P_2 T_2$$

Avogadro's Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Gay-Lussac's Law

Complete the following statements by filling in the blanks & writing the correct units.

As Pressure increases Volume decreases (increases) given that _____ is constant meaning that they have an inverse relationship.

As Volume increases Temperature increases (decreases) given that _____ is constant meaning that they have a direct inverse relationship.

As Pressure increases Temperature increases (decreases) given that _____ is constant meaning that they have a direct inverse relationship.

As the number of moles of gas in a container increases the Volume increases (decreases), the Pressure increases (decreases), and the Temperature increases (decreases).

Study Guide For Content Mastery Gas Laws

**Arch Wilkinson Shaw, Edward John
Mehren**



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