

STAAR GRADE 8 MATHEMATICS REFERENCE MATERIALS



LINEAR EQUATIONS

Slope-intercept form	$y = mx + b$
----------------------	--------------

Direct variation	$y = kx$
------------------	----------

Slope of a line	$m = \frac{y_2 - y_1}{x_2 - x_1}$
-----------------	-----------------------------------

CIRCUMFERENCE

Circle	$C = 2\pi r$	or	$C = \pi d$
--------	--------------	----	-------------

AREA

Triangle	$A = \frac{1}{2}bh$
----------	---------------------

Rectangle or parallelogram	$A = bh$
----------------------------	----------

Trapezoid	$A = \frac{1}{2}(b_1 + b_2)h$
-----------	-------------------------------

Circle	$A = \pi r^2$
--------	---------------

SURFACE AREA

	Lateral	Total
Prism	$S = Ph$	$S = Ph + 2B$
Cylinder	$S = 2\pi rh$	$S = 2\pi rh + 2\pi r^2$

VOLUME

Prism or cylinder	$V = Bh$
-------------------	----------

Pyramid or cone	$V = \frac{1}{3}Bh$
-----------------	---------------------

Sphere	$V = \frac{4}{3}\pi r^3$
--------	--------------------------

ADDITIONAL INFORMATION

Pythagorean theorem	$a^2 + b^2 = c^2$
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Simple interest	$I = Prt$
-----------------	-----------

Compound interest	$A = P(1 + r)^t$
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Staar Practice Exams 8th Math Texas

J Elliott



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Table of Contents Staar Practice Exams 8th Math Texas

1. Understanding the eBook Staar Practice Exams 8th Math Texas
 - The Rise of Digital Reading Staar Practice Exams 8th Math Texas
 - Advantages of eBooks Over Traditional Books
2. Identifying Staar Practice Exams 8th Math Texas
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Staar Practice Exams 8th Math Texas
 - User-Friendly Interface

4. Exploring eBook Recommendations from Staar Practice Exams 8th Math Texas
 - Personalized Recommendations
 - Staar Practice Exams 8th Math Texas User Reviews and Ratings
 - Staar Practice Exams 8th Math Texas and Bestseller Lists
5. Accessing Staar Practice Exams 8th Math Texas Free and Paid eBooks
 - Staar Practice Exams 8th Math Texas Public Domain eBooks
 - Staar Practice Exams 8th Math Texas eBook Subscription Services
 - Staar Practice Exams 8th Math Texas Budget-Friendly Options
6. Navigating Staar Practice Exams 8th Math Texas eBook Formats
 - ePub, PDF, MOBI, and More
 - Staar Practice Exams 8th Math Texas Compatibility with Devices
 - Staar Practice Exams 8th Math Texas Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Staar Practice Exams 8th Math Texas
 - Highlighting and Note-Taking Staar Practice Exams 8th Math Texas
 - Interactive Elements Staar Practice Exams 8th Math Texas
8. Staying Engaged with Staar Practice Exams 8th Math Texas
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Staar Practice Exams 8th Math Texas
9. Balancing eBooks and Physical Books Staar Practice Exams 8th Math Texas
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Staar Practice Exams 8th Math Texas
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Staar Practice Exams 8th Math Texas
 - Setting Reading Goals Staar Practice Exams 8th Math Texas
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Staar Practice Exams 8th Math Texas
 - Fact-Checking eBook Content of Staar Practice Exams 8th Math Texas
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ $2\text{NaOH}(\text{s}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{NaCl}(\text{s})$