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Student Exploration: Bohr Model of Hydrogen

Vocabulary: absorption spectrum, Bohr model, electron volt, emission spectrum, energy level, ionization energy, laser, orbital, photon

[Note to teachers and students: This Gizmo was designed as a follow-up to the Bohr Model: Introduction Gizmo. We recommend doing that activity before trying this one.]

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- What happens to an electron when it absorbs a tiny packet of light called a **photon**?
the electron absorbs the energy from it, causing a higher energy level
- What happens to an electron when it emits a photon?
the electrons lose energy, causing a lower energy level

Gizmo Warm-up

When white light is passed through hydrogen gas and separated by a prism, some wavelengths of light are absorbed by the hydrogen atoms in the gas. This causes black bars to appear in the **absorption spectrum** of hydrogen. An **emission spectrum** is created when hydrogen gas emits light.

The **Bohr model** of the atom was inspired by the spectrum produced by hydrogen gas. The link between light and atomic structure is illustrated by the *Bohr Model of Hydrogen* Gizmo. The Gizmo shows a **laser** pointed at a container of hydrogen gas. After passing through the gas, light from the laser goes through a prism and is detected on a screen.

- With the **Laser energy** set to 7.0 eV, click **Play** (). Observe the blue electron on the ORBITALS pane. Do 7.0-eV photons have any effect on the electron? No, it doesn't, the electron stays at lvl 1.
- Click **Pause** (). Set the **Laser energy** to 12.1 eV and click **Play**. What happens to the electron now? the electron shakes up and down and moves up by 2 orbital (level 3)
- Under **Go to energy level**, select **1**. Notice that a photon is emitted by the electron. What is the energy of the emitted photon? (12.1 eV)

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Table of Contents Student Exploration Bohr Model Of Hydrogen Answer Key

1. Understanding the eBook Student Exploration Bohr Model Of Hydrogen Answer Key
 - The Rise of Digital Reading Student Exploration Bohr Model Of Hydrogen Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Student Exploration Bohr Model Of Hydrogen Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Student Exploration Bohr Model Of Hydrogen Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Student Exploration Bohr Model Of Hydrogen Answer Key
 - Personalized Recommendations
 - Student Exploration Bohr Model Of Hydrogen Answer Key User Reviews and Ratings
 - Student Exploration Bohr Model Of Hydrogen Answer Key and Bestseller Lists
5. Accessing Student Exploration Bohr Model Of Hydrogen Answer Key Free and Paid eBooks

- Student Exploration Bohr Model Of Hydrogen Answer Key Public Domain eBooks
 - Student Exploration Bohr Model Of Hydrogen Answer Key eBook Subscription Services
 - Student Exploration Bohr Model Of Hydrogen Answer Key Budget-Friendly Options
6. Navigating Student Exploration Bohr Model Of Hydrogen Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Student Exploration Bohr Model Of Hydrogen Answer Key Compatibility with Devices
 - Student Exploration Bohr Model Of Hydrogen Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Student Exploration Bohr Model Of Hydrogen Answer Key
 - Highlighting and Note-Taking Student Exploration Bohr Model Of Hydrogen Answer Key
 - Interactive Elements Student Exploration Bohr Model Of Hydrogen Answer Key
8. Staying Engaged with Student Exploration Bohr Model Of Hydrogen Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Student Exploration Bohr Model Of Hydrogen Answer Key
9. Balancing eBooks and Physical Books Student Exploration Bohr Model Of Hydrogen Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Student Exploration Bohr Model Of Hydrogen Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Student Exploration Bohr Model Of Hydrogen Answer Key
- Setting Reading Goals Student Exploration Bohr Model Of Hydrogen Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Student Exploration Bohr Model Of Hydrogen Answer Key
- Fact-Checking eBook Content of Student Exploration Bohr Model Of Hydrogen Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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