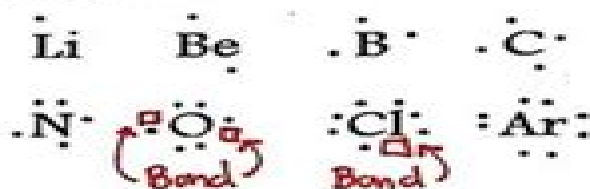


Lewis (Electron) Dot Diagrams:

Below you will find eight (8) electron dot diagrams for common elements. The dots around the element's symbol represent a certain type of electrons.



- What type of electrons possessed by an atom in its outermost energy level do the dots represent?

Valence electrons

- Why do you think these electrons in Carbon, Nitrogen, and Oxygen especially are spread out in the way that they are?

e^- repel until they have no choice

- If Chlorine has 7 dots, how many more can it possibly gain? 1 $7+1=8$
 - If Oxygen has 6 dots, how many more can it possibly gain? 2 $6+2=8$
 - What is the correlation between open electron spaces and possible bonds formed?

Direct - open space is a location for a bond

Practice:

1. Write the electron dot diagram for each of the following elements:

a. Hydrogen -
1 ve.



b. Barium -
2 ve.



c. Krypton -
8 ve.



d. Magnesium -
2 ve.



e. Oxygen -
6 ve.



f. Chlorine -
7 ve.



* Often, the .. will orient next to each other

2. Use electron dot diagrams to show the transfer of electrons in each of the following ionic bonds. Then, write the formula for each combination - Cation(s) first!

a. ^MPotassium and Chlorine



b. Beryllium and Fluorine



c. Rubidium and Oxygen



d. Calcium and Sulfur



Unit 5 Chemical Bonding Test Answers

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