

Worksheet 1 – Scientific Notation and Significant Figures

1. Convert each of the following into scientific notation.

- | | | |
|---------------------------|-------------------------|--------------------------|
| a) 3427 | b) 0.00456 | c) 123,453 |
| d) 172 | e) 0.000984 | f) 0.502 |
| g) 3100.0×10^2 | h) 0.0114×10^4 | i) 107.2 |
| j) 0.0000455 | k) 2205.2 | l) 30.0×10^{-2} |
| m) 0.982×10^{-3} | n) 0.0473 | o) 650.502 |
| p) 3.03×10^{-1} | q) 20.4×10^5 | r) 1.29 |
| s) 0.00565 | t) 1362205.2 | u) 450.0×10^3 |
| v) 1000×10^{-3} | | |

2. Determine the number of significant figures in each of the following:

- | | | |
|---------------------------|-------------------------|--------------------------|
| a) 3427 | b) 0.00456 | c) 123,453 |
| d) 172 | e) 0.000984 | f) 0.502 |
| g) 3100.0×10^2 | h) 0.0114×10^4 | i) 107.2 |
| j) 0.0000455 | k) 2205.2 | l) 30.0×10^{-2} |
| m) 0.982×10^{-3} | n) 0.0473 | o) 650.502 |
| p) 3.03×10^{-1} | q) 20.4×10^5 | r) 1.29 |
| s) 0.00565 | t) 1362205.2 | u) 450.0×10^3 |
| v) 1000×10^{-3} | w) $546,000 \pm 10$ | x) $546,000 \pm 1000$ |

3. Convert each into decimal form.

- | | | | |
|-----------------------|---------------------------|------------------------|---------------------|
| 1.56×10^4 | 0.56×10^{-2} | 3.69×10^{-2} | 736.9×10^5 |
| 0.00259×10^5 | 0.000459×10^{-1} | 13.69×10^{-2} | 6.9×10^4 |
| 0.00259×10^3 | 0.0209×10^{-3} | | |

4. Calculate the following. Give the answer in correct scientific notation.

- | | | | |
|--|-----------------------------------|---|--|
| a) 4.53×10^5
+ 2.2×10^6 | b) 1913.0
- 4.6×10^3 | c) 2.34×10^{-4}
+ 1.92×10^{-23} | d) 2.130×10^3
- 6.6×10^2 |
| e) 9.10×10^3
+ 2.2×10^6 | f) 1113.0
- 14.6×10^2 | g) 6.18×10^{-45}
+ 4.72×10^{-44} | h) 4.25×10^{-3}
- 1.6×10^{-2} |

5. Calculate the following. Give the answer in correct scientific notation.

- a) $3.95 \times 10^2 / 1.5 \times 10^6$ b) $(3.5 \times 10^2)(6.45 \times 10^{10})$ c) $4.44 \times 10^7 / 2.25 \times 10^5$ d) $(4.50 \times 10^{-12})(3.67 \times 10^{-12})$
- e) $1.05 \times 10^{-26} / 4.2 \times 10^{-26}$ f) $(2.5 \times 10^9)(6.45 \times 10^4)$ g) $6.022 \times 10^{23} / 3.011 \times 10^{-26}$ h) $(6.68 \times 10^5)(3.45 \times 10^{-10})$

Unit 08 Practice Problems Chemistry Key Significant Figures

Charles Henrickson



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Dasgupta, Kevin A. Schug, 2020 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry Comprehensive Chemistry XI Dr. B. Kapila, S. K. Khanna, 2010-11 Comprehensive chemistry according to the new syllabus prescribed by Central Board of Secondary Education CBSE Investigating Chemistry Matthew Johll, 2008-12-22 In its new second edition Investigating Chemistry A Forensic Science Perspective remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package **Ebook: Chemistry: The Molecular Nature of Matter and Change** Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change *Chemistry: Core Concepts, 3rd Edition* Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package *Chemistry* Carson-Dellosa Publishing, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards **Foundations of College Chemistry, Alternate** Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped

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