

UNIT 3

CLASSIFICATION OF ELEMENTS & PERIODICITY IN PROPERTIES

A well organized & tabulated classification of elements help to locate, identify & characterize the elements & its properties.

VARIOUS ATTEMPTS OF THE CLASSIFICATION

DOBEREINER'S TRIADS

Element	Atomic weight	Element	Atomic weight	Element	Atomic weight
Li	7	Ca	40	Cl	35.5
Na	23	Sr	88	Br	80
K	39	Ba	137	I	127

Main Features

- ✓ Dobereiner classified the elements into groups of three elements (called **triads**) with similar properties.
 - ✓ the atomic weight of the middle element was the arithmetic mean of the other two, e.g.
Mean of atomic weight of *Na* = $\frac{7(\text{atomic weight of Li}) + 39(\text{atomic weight of K})}{2} = 23$
- This Dobereiner's relationship is referred as "**Law of triads**".

Limitations

- ✓ Dobereiner could not arrange all the known elements into triads.
- ✓ Dobereiner could identify only three such triads.

NEWLAND'S OCTAVES

Element	Li	Be	B	C	N	O	F
Atomic weight	7	9	11	12	14	16	19
Element	Na	Mg	Al	Si	P	S	Cl
Atomic weight	23	24	27	29	31	32	35.5

Main Features

- ✓ Newland arranged the elements in the increasing order of atomic weights.
- ✓ The properties of the every eighth element are similar to the first one, e.g.

Properties of *Li* are same as *Na* (the eighth element to Li)

This relationship is called as "**Law of octaves**".

Limitations

- ✓ This classification was successful up to the element Ca.
- ✓ When Noble gas elements were discovered at a later stage, their inclusion in these octaves disturbed the entire arrangement.

MENDELEEV'S PERIODIC TABLE

Main Features

- ✓ The Chemical & physical properties of elements are a periodic function of their atomic masses.
- ✓ This relationship is called as "**Mendeleev's Periodic Law**".
- ✓ Elements are arranged in tabular form in Rows & columns.

Unit 27 Chemical Periodicity Its Applications

G Kelville Davis



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