

Physical and Chemical Changes Lab

Name _____

25 Points

Period _____ Date _____

Introduction: As you read this it is probably fall. Summer flowers are fading and dying. Leaves are changing from green to red, yellow and orange. All these changes involve chemistry.

As you have learned, chemistry is the study of matter and the changes that it undergoes. These changes can be classified as physical, chemical, or nuclear. This lab will focus on physical and chemical changes. When a physical change occurs, the physical properties of a substance – such as its size, shape, state, or density – are altered, but its chemical composition remains the same. Examples of physical changes include melting ice, crushing gravel, tearing paper, grinding pepper, and boiling water. No new substances are formed as a result of these changes.

Chemical changes, also known as chemical reactions, result in the formation of one or more new substances with different chemical properties and compositions from the original material. Examples of chemical changes include plants dying, leaves changing color, paper burning, bananas ripening, bread baking, or iron rusting. Some signs of chemical changes include a change in color, the formation of a precipitate (a new solid substance that settles out of solution), the production and release of a gas, or a change in temperature.

In this investigation, you will conduct tests on several substances and then use your data to determine whether the resulting changes were chemical or physical. As you observe each change, remember to ask yourself, "Has the change altered the identity of the substance?"

Pre-Lab Questions:

1. What are the 4 observable signs that indicate a chemical change is taking place?

2. Define precipitate as it relates to this lab. _____

3. Identify the following as either a chemical (C) or physical (P) change:

- | | |
|------------------------------|-------|
| a. burning wood | _____ |
| b. dry ice changing to a gas | _____ |
| c. freezing water | _____ |
| d. ripening fruit | _____ |
| e. dissolving sugar in water | _____ |

4. Why are you instructed to feel outside the test tube after two chemicals are mixed, as in Steps 5 and 6 of this investigation? _____

5. What are some safety cautions you need to observe during this investigation? _____

Study Of Chemical Changes Lab Answer Key

Valentine Korah



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