

Thermochemistry Formula Sheet:

1st Law of Thermodynamics: Endothermic: $q = +\cancel{L}$ Exothermic: $q = -\cancel{L}$	Internal Energy: $\Delta U = q + W$
Work Done By/On a Gas: Work done by the system: $W = -\cancel{L}$ Work done on the system: $W = +\cancel{L}$	$W = -P \Delta V$ $1 L \cdot atm = 101.3 \text{ Joules}$
Specific Heat Capacity: $J/(g \cancel{L} ^\circ C) \cancel{L}$	$C_s = \frac{q}{m \Delta T}$
Molar Heat Capacity: $J/(mol ^\circ C)$	$C_M = \frac{q}{n \Delta T}$
Heat Capacity: $J/^\circ C$	$C_n = \frac{q}{\Delta T}$
Calorimetry - Temperature Change:	$q = mc \Delta T$ $C_{H_2O} = 4.184 \frac{J}{g ^\circ C} \quad C_{ice} = 2.09 \frac{J}{g ^\circ C} \quad C_{steam} = 2.03 \frac{J}{g ^\circ C}$
Calorimetry - Phase Change:	$q = m \Delta H \quad q = n \Delta H$ $H_2O: \Delta H_{fus} = 334 J/g \quad \Delta H_{fus} = 6.01 kJ/mol$ $H_2O: \Delta H_{vap} = 2260 J/g \quad \Delta H_{vap} = 40.7 kJ/mol$
Final Temperature of a Mixture: $T_A \rightarrow$ Initial Temp. of Substance A $T_B \rightarrow$ Initial Temp. of Substance B	$-q_A = q_B$ $-m_A C_A (T_f - T_A) = m_B C_B (T_f - T_B)$ Note: This equation will work if there are no phase changes.
Coffee Cup Calorimeter:	$\Delta H^\circ_{rxn} = \frac{q_{rxn}}{n} \quad q_{rxn} = -q_{H_2O}$
Enthalpy Change:	

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Table of Contents Thermochemistry Cp Test Sheet

1. Understanding the eBook Thermochemistry Cp Test Sheet
 - The Rise of Digital Reading Thermochemistry Cp Test Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Cp Test Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Cp Test Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Cp Test Sheet
 - Personalized Recommendations
 - Thermochemistry Cp Test Sheet User Reviews and Ratings
 - Thermochemistry Cp Test Sheet and Bestseller Lists
5. Accessing Thermochemistry Cp Test Sheet Free and Paid eBooks

- Thermochemistry Cp Test Sheet Public Domain eBooks
 - Thermochemistry Cp Test Sheet eBook Subscription Services
 - Thermochemistry Cp Test Sheet Budget-Friendly Options
6. Navigating Thermochemistry Cp Test Sheet eBook Formats
- ePub, PDF, MOBI, and More
 - Thermochemistry Cp Test Sheet Compatibility with Devices
 - Thermochemistry Cp Test Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Thermochemistry Cp Test Sheet
 - Highlighting and Note-Taking Thermochemistry Cp Test Sheet
 - Interactive Elements Thermochemistry Cp Test Sheet
8. Staying Engaged with Thermochemistry Cp Test Sheet
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Cp Test Sheet
9. Balancing eBooks and Physical Books Thermochemistry Cp Test Sheet
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Cp Test Sheet
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermochemistry Cp Test Sheet
- Setting Reading Goals Thermochemistry Cp Test Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermochemistry Cp Test Sheet
- Fact-Checking eBook Content of Thermochemistry Cp Test Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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