

Thermochemistry Problems - Worksheet Number Two (answers available on web site)

The older energy unit of calories has not been discussed in class. You may see it from time to time. The conversion is $1.000 \text{ cal} = 4.184 \text{ J}$. All the calculation techniques are the same regardless of using calories or Joules.

- Convert from one unit to the other:
 - 1.69 Joules to calories
 - 0.3587 J to cal
 - 820.1 J to kilocalories
 - 68 calories to kilocalories
 - 423 calories to kilocalories
 - 20.0 calories to Joules
 - 252 cal to J
 - 2.45 kilocalories to calories
 - 556 kilocalories to cal
 - 6.78 kilocalories to kilojoules
 - 59.6 calories to kcal
 - 449.6 joules to kilojoules
 - 9.806 kJ to J
 - 5.567 cal to J
 - 5467.9 kcal to J
- Determine the temperature change when:
 - 20.0 g of water is heated from 16.8°C to 39.2°C .
 - 35.0 g of water is cooled from 56.5°C to 5.9°C .
 - 50.0 g of liquid water is heated from 0.0°C to 100.0°C .
 - 25.0 g of ice is warmed from -25.0°C to 0.0°C , but does not melt.
 - 30.0 g of steam heats from 373.2 K to 405.0 K .
- Determine the energy required (in Joules) when the temperature of 3.21 grams of water increases by 4.0°C while remaining liquid.
- Determine the energy needed (in Joules) when 55.6 grams of water at 43.2°C is heated to 78.1°C .
- Determine the energy required (in kilojoules) when cooling 456.2 grams of water at 89.2°C to a final temperature of 5.9°C .
- Determine the energy required to:
 - melt 5.62 moles of ice at 0°C .
 - melt 74.5 grams of ice at 0°C .
 - boil 0.345 moles of water at 100.0°C .
 - boil 43.89 grams of water at 100.0°C .
- Determine the energy change involved to:
 - Convert 16.2 grams of ice to liquid water.
 - Convert 5.8 grams of water to steam
 - Convert 98.2 grams of water to ice.
 - Convert 52.6 grams of steam to water
 - Convert 34.0 grams of water at 20.0°C to steam at 100.0°C .
 - Convert 125.0 grams of ice at 0.0°C to steam at 100.0°C .
 - Convert 25.9 grams of steam at 100.0°C to ice at 0.0°C .

Thermochemistry Problems Worksheet Number Two

Daniel C. Tofan



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