

Thermochemistry Quiz 1

Questions 1-4 (1 pt. each); Questions 5-6 (2 pts. Each)

- _____ 1. An endothermic reaction causes the surroundings to
A. warm up.
B. become acidic.
C. condense.
D. decrease in temperature.
E. release CO₂
- _____ 2. According to the first law of thermodynamics:
A. energy is neither lost nor gained in any energy transformations.
B. perpetual motion is possible.
C. energy is conserved in quality but not in quantity.
D. energy is being created as time passes. We have more energy in this universe now than when time began.
- _____ 3. A calorimeter
A. is equal to the molar enthalpy of a reaction.
B. is a dieting aid.
C. is an indicator of a spontaneous reaction.
D. is a device used to measure the transfer of heat energy.
E. is useful in measuring the amount of heat released by endothermic reactions.
- _____ 4. The specific heat of gold is 0.129 J/g•°C. What is the molar heat capacity of gold?
A. 0.039 J/mol•°C
B. 0.129 J/mol•°C
C. 25.4 J/mol•°C
D. 39.0 kJ/mol•°C
E. 197 J/mol•°C
- _____ 5. Copper metal has a specific heat of 0.385 J/g•°C. Calculate the amount of heat required to raise the temperature of 22.8 g of Cu from 20.0°C to 875°C.
A. 1.97×10^{-2} J
B. 1.0×10^{-2} J
C. 329 J
D. 7.51 kJ
E. 10.5 kJ
- _____ 6. The combustion of butane produces according to the equation
$$2\text{C}_4\text{H}_{10}(\text{g}) + 13\text{O}_2(\text{g}) \rightarrow 8\text{CO}_2(\text{g}) + 10\text{H}_2\text{O}(\text{l}) \quad \Delta H^\circ_{\text{rxn}} = -5,314 \text{ kJ}$$

How many grams of butane must be burned to release 6,375 kJ of heat?
A. 1.20 g
B. 139 g
C. 0.0413 g
D. 69.7 g
E. 97.8 g

Questions 7-8: Label the following processes as either *exothermic* or *endothermic*. (1 pt. each)

7. Dry ice (solid CO₂) sits at room temperature and sublimates to CO₂ gas. _____
8. A block of solid aluminum is heated until it melts. _____

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