Thermochemistry Problems

The First Law

- The complete combustion of acetic acid, HC2H3O28, to form water, H2O88, and CO288, at constant pressure releases 871.7 kJ of heat per mol of acetic acid. Write a balanced thermochemical equation for this reaction. Draw an enthalpy diagram for the reaction.
- 2. Consider the following reaction, which occurs at standard state:

 $2Cl_{10} \rightarrow Cl_{20}$ $\Delta H^{\circ} = -243.4 \text{ kJ}$

- a. Which has the higher, more positive, enthalpy under these conditions, 2Cl to or Cl2to?
- b. Consider the following reaction: 2 Mg_(i) + O_{2(j)} → 2 MgO_(i) ΔH = -1204 kJ.
- c. Is the reaction exothermic or endothermic?
- d. Calculate the amount of heat transferred when 2.4 g of Mg_(c) reacts at constant pressure. How many grams of MgO are produced during an enthalpy change of 96.0 kJ?
- e. How many kilo joules of heat are absorbed when 7.50 g of MgO_{i0} are decomposed into Mg_{i0} and O2_{i0} at constant pressure?

Enthalpy

Calorimetry Problems

- Determine the specific heat of a sample of Cu from the fact that 64.0J are needed to raise the temperature of 15.0 g of Cu metal from 22.0°C to 33.0°C. [0.388 J/g°C]
- 4. The specific heat of a sample of copper is 0.385 J/g-K. How many joules of heat are necessary to raise the temperature of a 1.42 kg block of copper from 25.0 °C to 88.5°C?
- A 50.0-g sample of water at 100.00°C was placed in an insulated cup. Then 25.3-g
 of zinc at 25.00°C was added to the water. The temperature of the water
 dropped to 96.68°C. What is the specific heat of the zinc?[0.388 J/g°C]
- When a 6.50 g sample of solid NaOH dissolves in 100.0g of water in a coffee-cup calorimeter, the temperature rises from 21.6°C to 37.8°C. Calculate the ΔH (in kJ/mol NaOH) for the solution process. (Assume the specific heat of the solution formed is 4.18 J/g-K)

 $NaOH_{(a)} \rightarrow Na^{*}(aq)+OH^{*}(aq)$

A house is designed to have passive solar energy features, Brickwork is to be incorporated into the interior of the house to act as a heat absorber. Each brick

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