

THERMODYNAMICS TEST REVIEW ANSWERS

Specific Heat Problems

1. How many J would it take to raise the temperature of 200 grams of water from 5 degree Celsius C to 85 degree Celsius?

Answer: $Q = (200 \text{ g})(4.18)(80 \text{ }^{\circ}\text{C}) = 66880 \text{ J}$

2. How many J would problem number 1 be if it was aluminum instead of water?

Answer: $Q = (200 \text{ g})(0.897)(80 \text{ }^{\circ}\text{C}) = 14352 \text{ J}$

3. How many grams of copper could be heated from 20 °C to 75 °C if 1200J are applied to it?

Answer: $1200 \text{ J} = (x)(0.385)(55 \text{ }^{\circ}\text{C}) = 56.7 \text{ g}$

4. What is the specific heat capacity of a substance if 750 J caused 100 grams of it to go from 90 °C to 135 °C?

Answer: $750 \text{ J} = (100 \text{ g})(x)(45 \text{ }^{\circ}\text{C}) = 0.167 \text{ J/g}^{\circ}\text{C}$

5. What would the final temperature be if 500 J are applied to 150 grams of ice at -90 °C?

Answer: $500 \text{ J} = (150 \text{ g})(2.06)(x - (-90 \text{ }^{\circ}\text{C})) = -88.4 \text{ }^{\circ}\text{C}$

6. What would the temperature change by if a 90 gram piece of hot iron cooled by losing 200 J.

Answer: $-200 \text{ J} = (90 \text{ g})(0.449)(x) = -4.95 \text{ }^{\circ}\text{C}$

7. What was the initial temperature if 500 J were applied to 250 g of mercury and the final temperature was 50 °C?

Answer: $500 \text{ J} = (250 \text{ g})(0.140)(50 \text{ }^{\circ}\text{C} - x) = 35.7 \text{ }^{\circ}\text{C}$

Calorimetry

1. What is the specific heat capacity of water?

Answer: $4.18 \text{ J/g}^{\circ}\text{C}$

2. A piece of ice is placed on the sidewalk on a hot summer day. Describe what happens to energy content of the piece of ice, does it gain or lose energy? What happens to the air around the ice, does it gain or lose energy?

Answer: The ice gains energy from the surrounding air and increases its energy (endothermic). The air loses energy to the ice and decreases its energy (exothermic)

Thermodynamics Test Review Answers

The Princeton Review



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Table of Contents Thermodynamics Test Review Answers

1. Understanding the eBook Thermodynamics Test Review Answers
 - The Rise of Digital Reading Thermodynamics Test Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamics Test Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Test Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermodynamics Test Review Answers
 - Personalized Recommendations
 - Thermodynamics Test Review Answers User Reviews and Ratings
 - Thermodynamics Test Review Answers and Bestseller Lists
5. Accessing Thermodynamics Test Review Answers Free and Paid eBooks
 - Thermodynamics Test Review Answers Public Domain eBooks
 - Thermodynamics Test Review Answers eBook Subscription Services
 - Thermodynamics Test Review Answers Budget-Friendly Options
6. Navigating Thermodynamics Test Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermodynamics Test Review Answers Compatibility with Devices
 - Thermodynamics Test Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermodynamics Test Review Answers
 - Highlighting and Note-Taking Thermodynamics Test Review Answers
 - Interactive Elements Thermodynamics Test Review Answers
8. Staying Engaged with Thermodynamics Test Review Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Thermodynamics Test Review Answers
- 9. Balancing eBooks and Physical Books Thermodynamics Test Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Test Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermodynamics Test Review Answers
 - Setting Reading Goals Thermodynamics Test Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermodynamics Test Review Answers
 - Fact-Checking eBook Content of Thermodynamics Test Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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