

# How To Find Theoretical Yield

Theoretical yield is the mass of product that the limiting reactant produces. All we need to do is to convert moles of the product that the limiting reactant gives to grams.

*Example: If 10 grams of Zn reacted with 20 grams of HCl, what is the theoretical yield of hydrogen gas?*



Transformation  
TUTORING

Zn:

limiting

$$10\text{g Zn} \times \frac{1\text{mol Zn}}{65.38\text{g Zn}} \times \frac{1\text{mol H}_2}{1\text{mol Zn}} = 0.153\text{mol H}_2$$

molar mass Zn      coefficients from equation

SMALLER

$$\text{HCl} : 20\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} \times \frac{1\text{mol H}_2}{2\text{mol HCl}} = 0.274\text{mol H}_2$$

$$0.153\text{mol H}_2 \times \frac{2.016\text{g}}{1\text{mol H}_2} = 0.308\text{g H}_2$$

molar mass H<sub>2</sub>

theoretical yield of hydrogen gas

# Theoretical Yield Limiting Reagents Answer Key

**Julia Burdge**



## **Theoretical Yield Limiting Reagents Answer Key:**

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