

Thermochemistry Practice Problems

KEY

Flow of Energy:

1. How many kJ of energy are in a donut that contains 200.0 Calories?

$$\frac{200 \text{ Cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ Cal}} = \boxed{836.8 \text{ kJ}}$$

2. What is the specific heat of a substance that has a mass of 25.0 g and requires 525.0 calories to raise its temperature by 15.0°C?

$$q = mc\Delta T$$

$$m = 25.0 \text{ g}$$

$$q = 525.0 \text{ cal}$$

$$\Delta T = 15.0^\circ\text{C}$$

$$525.0 \text{ cal} = (25.0 \text{ g})(c)(15.0^\circ\text{C})$$

$$\boxed{c = 1.40 \text{ cal/g}^\circ\text{C}}$$

3. Suppose 100.0 g of H₂O(s) absorbs 1255.0 J of heat. What is the corresponding temperature change?

The specific heat capacity of H₂O(s) is 2.1 J/g°C. $q = mc\Delta T$

$$q = 1255.0 \text{ J}$$

$$m = 100.0 \text{ g}$$

$$c = 2.1 \text{ J/g}^\circ\text{C}$$

$$1255.0 \text{ J} = (100.0 \text{ g})(2.1 \text{ J/g}^\circ\text{C})(\Delta T)$$

$$\boxed{\Delta T = 6.0^\circ\text{C}}$$

4. How many joules of heat energy are required to raise the temperature of 100.0 g of aluminum by 120.0°C? The specific heat capacity of aluminum is 0.90 J/g°C.

$$q = mc\Delta T$$

$$m = 100.0 \text{ g}$$

$$c = 0.90 \text{ J/g}^\circ\text{C}$$

$$\Delta T = 120.0^\circ\text{C}$$

$$q = (100.0 \text{ g})(0.90 \text{ J/g}^\circ\text{C})(120.0^\circ\text{C})$$

$$\boxed{q = 11,000 \text{ J}}$$

5. In reference to #4, what is the molar heat capacity of aluminum?

$$(0.90 \text{ J/g}^\circ\text{C})(26.98 \text{ g/mol}) = \boxed{24 \text{ J/mol}^\circ\text{C}}$$

Measuring and Expressing Heat Changes:

6. A student mixed 75.0 mL of water containing 0.75 mol HCl at 25°C with 75.0 mL of water containing 0.75 mol of NaOH at 25°C in a foam cup calorimeter. The temperature of the resulting solution increased to 35°C. How much heat in kilojoules was released by this reaction ($C_{\text{water}} = 4.18 \text{ J/g}^\circ\text{C}$)?

$$q = mc\Delta T$$

$$m = 150.0 \text{ g}$$

$$c = 4.184 \text{ J/g}^\circ\text{C}$$

$$\Delta T = 10^\circ\text{C}$$

$$q = (150.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(10^\circ\text{C})$$

$$\boxed{q = 6300 \text{ J}}$$

7. Calculate the amount of heat evolved when 150 g of water are reacted with excess calcium according to the following thermochemical equation:



$$\frac{150 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}} \times \frac{-65.2 \text{ kJ}}{1 \text{ mol H}_2\text{O}} = \boxed{-270 \text{ kJ}; \text{ exothermic}}$$

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ACT



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