

Thermodynamic Feasible Reactions: (2000-2019)

10. Calculate the standard enthalpy of formation for the compound



$$\Delta H^\circ = 2^\circ$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$

$$\Delta H^\circ = 210.9 \text{ kJ}$$

11. Calculate the free enthalpy of the combination of weight percent: ΔH°



$$\Delta H^\circ = 2^\circ$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$

$$\Delta H^\circ = 210.9 \text{ kJ}$$

12. Calculate the total required for thermodynamic stability: CaCO_3 , BaCO_3 , CaCO_3 , and BaCO_3 stability.



$$\Delta H^\circ = 210.9 \text{ kJ}$$

13. $\text{CaCO}_3 \rightarrow \Delta \text{CaCO}_3 \rightarrow \Delta \text{CaCO}_3 \rightarrow \Delta \text{CaCO}_3$ $\Delta H^\circ = 2^\circ$



$$\Delta H^\circ = 210.9 \text{ kJ}$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$

$$\Delta H^\circ = 210.9 \text{ kJ}$$

14. Calculate the total required for thermodynamic stability: CaCO_3 , BaCO_3 , CaCO_3 , and BaCO_3 stability.



$$\Delta H^\circ = 210.9 \text{ kJ}$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$



$$\Delta H^\circ = 210.9 \text{ kJ}$$

$$\Delta H^\circ = 210.9 \text{ kJ}$$

15. The thermodynamic stability change is related to the thermodynamic stability.



$$\Delta H^\circ = 210.9 \text{ kJ}$$

16. The thermodynamic stability change is related to the thermodynamic stability.



$$\Delta H^\circ = 210.9 \text{ kJ}$$

Thermochemistry Practice Problems Answer Key

R Bogdan



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