These problems are not meant to introduce the problems associated with thermochemistry. You should already have been introduced to the concepts in your lecture. These are just problems for extra practice. They are presented in random order, on purpose, to let you practice RECOGNIZING what TYPE of problem is in front of you, so that you know how to approach the problem.

Often you will find it much easier when you can identify the type of problem first.

1. From the following enthalpies of reaction:

$H_2(g) + F_2(g) \longrightarrow 2HF(g)$	$\Delta H = -537 \text{ kJ}$
$C(s) + 2F_2(g) \longrightarrow CF_4(g)$	$\Delta H = -680 \text{ kJ}$
$2C(s) + 2H_2(g) \longrightarrow C_2H_4(g)$	$\Delta H = +52.3 \text{ kJ}$
calculate AH for the reaction of eth	whene with Ev-

 $C_2H_4(g) + 6F_2(g) \longrightarrow 2CF_4(g) + 4HF(g)$

- When a 4.25-g sample of solid ammonium nitrate dissolves in 60.0 g of water in a coffee-cup calorimeter, the temperature drops from 22.0 °C to 16.9°C. Calculate ΔH in (kJ/mol NH₄NO₃) for the solution process NH₄NO₃ (s) → NH₄* (aq) + NO₃* (aq)
 Assume that the specific heat of the solution is the same as that of pure water.
- CH₃OH (g) → CO (g) + 2H₂ (g) ΔH = +90.7 kJ
 Calculate the amount of heat transferred when 45.0 g of CH₃OH (g) are decomposed by this reaction.
- substance AHA kJ/mod Calculate the standard enthalpy change for the following reactions: SiCL (D) -640.1 $SiCl_4(l) + 2H_2O(l) \longrightarrow SiO_2(s) + 4HCl(g)$ -910.9SiO₂ (s) HCI (aa) -167.2HCI (g) -92.30H₂O (g) -241.825. $2\text{Mg}(s) + \text{O}_2(g) \longrightarrow 2\text{MgO}(s)$ $\Delta H = -1204 \text{ kJ}$ H-O (I) -285.83
- How many kJ of heat are absorbed when 7.50 g of MgO (s) are decomposed into Mg (s) & O₂ (g)?
- Write a balanced equation that corresponds to ΔH_i° for SO₂ (g).
- A 440. g quantity of a certain metal is heated to 100.0 °C. It is immediately thrust into 258 g of
 water that is initially at 25.0 °C. The temperature of the metal-water mixture rises to 36.5 °C. What
 is the specific heat of the metal? Assume the density of the water is 0.998 g /mL and the specific
 heat capacity of the water is 4.184 J/g. °C.

Thermochemistry Practice Test A Answers

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