Thermochemistry Worksheet #1

AP Chemistry, Chapter 6

ANSWERS

- +1000 J, endothermic (q is positive)
- 2) -2267 J, work done by system on the surroundings
- 3) $\Delta E = +300 \text{ J}, q = -200 \text{ J exothermic}$
- 0.5354 kJ, 0.1280 kcal
- 5) q = +4184 J; at constant P, so ΔH = q or ΔH = +4184 J; w = -30.4 J; ΔE = 4154 J
- 6) q = +6500 J, w = −9620 J, ∆E = −3120 J
- 7) a. endothermic, 1.8 kJ = 0.43 kcal
 - b. exothermic, -2.4 kcal = -10 kJ
 - c. exothermic, -0.025 kcal = -0.10 kJ
 - d. exothermic, -95.3 kJ = -22.8 kcal
- 8) 0.0329 J/g°C or 0.00786 cal/g°C
- 9) 59.7 kJ
- 10) -49.4 kJ/mol, exothermic
- 11) –289 kJ/mol H₂
- 12) 18.9 L
- 13) -5670 kJ/mol sucrose, -11200 kJ/mol stearic acid
- 14) $\Delta H_{rxn} = -822.3 \text{ kJ}$
- 15) $\Delta H_{rm} = -78.7 \text{ kJ}$

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