

## Specific Heat Worksheet

$C_p = q/m\Delta T$ , where  $q$  = heat energy,  $m$  = mass, and  $T$  = temperature

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the heat capacity of iron.

$$1086.75 = 15.75g \cdot 150 \cdot C_p$$

$$1086.75 = 2362.5 C_p$$

$$C_p = .46 \text{ J/g}^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$q = C_p \cdot m \cdot \Delta T = .90(10)(33)$$

$$= 297 \text{ J}$$

$$\frac{55}{33}$$

3. To what temperature will a 50.0 g piece of glass raise if it absorbs 5275 joules of heat and its heat capacity is 0.50 J/g°C? The initial temperature of the glass is 20.0°C.

$$5275 = .50(50)\Delta T$$

$$\Delta T = 211$$

$$T_f = \frac{211}{20} = 231$$

$$231^\circ\text{C}$$

4. Calculate the heat capacity of a piece of wood if 1500.0 g of the wood absorbs  $6.75 \times 10^4$  joules of heat, and its temperature changes from 32°C to 57°C.

$$67500 = C_p \cdot 1500 \cdot 25^\circ\text{C}$$

$$67500 = 37500 C_p$$

$$1.8 \text{ J/g}^\circ\text{C}$$

5. 100.0 mL of 4°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat needed to cause this rise in temperature.

$$Q = 100g \cdot 4.18 \cdot 33^\circ\text{C}$$

$$= 13794$$

$$13794 \text{ J}$$

$$13794$$

6. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the heat capacity of mercury.

$$455 = C_p \cdot 25g \cdot 130$$

$$455 = 3250 C_p$$

$$.14 = C_p$$

7. What is the Specific heat capacity of silver metal if 55.00 g of the metal absorbs 198 joules of heat and the temperature rises 15.0°C?

$$198 = C_p \cdot 55g \cdot 15^\circ\text{C}$$

$$198 = 825 C_p$$

$$C_p = .24 \text{ J/g}^\circ\text{C}$$

8. What is the change in temperature of 150.0 g of chloroform if it absorbs 1000.0 joules of heat, and the specific heat of chloroform is 0.96 J/g°C?

$$1000 = .96(150g)\Delta T$$

$$1000 = 144 \Delta T$$

$$6.94 = \Delta T$$

$$6.9^\circ\text{C}$$

## Answers

- $C_p = .46 \text{ J/g}^\circ\text{C}$
- $Q = 297 \text{ J}$
- $T_{\text{final}} = 231^\circ\text{C}$
- $C_p = 1.8 \text{ J/g}^\circ\text{C}$
- $Q = 13800 \text{ J}$   $13794$
- $C_p = .14 \text{ J/g}^\circ\text{C}$
- $C_p = .24 \text{ J/g}^\circ\text{C}$
- $\Delta T = 6.9^\circ\text{C}$

# Thermochemistry And Specific Heat Problems Canswers

**H Kauffman**



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