

Answers - Exercises (Thermochemistry)

Q1: Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15g of copper from 25°C to 60°C. (Ans : $c = 0.39 \text{ J/g.K}$)

Answer:

$$q = mc\Delta T$$

$$q = 204.75 \text{ J}$$

$$m = 15\text{g}$$

$$\Delta T = (60 - 25) = 35^\circ\text{C}$$

$$c = \frac{q}{m\Delta T}$$

$$= \frac{204.75 \text{ J}}{(15\text{g})(35^\circ\text{C})}$$

$$= 0.39 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$$

Q2: A chemist burns 0.8650 g of graphite (a form of carbon) in a new bomb calorimeter, and CO_2 forms. If 393.5 kJ of heat is released per mole of graphite and T increases 2.613 K, what is the capacity of the bomb calorimeter? (Ans: 10.84 kJ/K)

Answer:



$$\text{Mass of Carbon (Graphite)} = 0.8650 \text{ g}$$

$$\text{Atomic mass of Carbon} = 12 \text{ g/mol}$$

$$\text{Mol of graphite} = \text{mass/atomic mass} = 0.8650 \text{ g/} 12\text{g mol}^{-1} = 0.072 \text{ mol}$$

$$\text{Heat of combustion, } \Delta H_{\text{comb}} = -393.5 \text{ kJ/mol}$$

$$\text{Temperature change, } \Delta T = 2.613 \text{ K}$$

$$\text{Calorimeter heat capacity, } C_{\text{cal}} = ?$$

$$\Delta H_{\text{comb}} = \frac{q_{\text{rxn}}}{n} = \frac{C_{\text{cal}} \times \Delta T}{n}$$

$$393.5 = \frac{C_{\text{cal}}(2.613)}{0.072}$$

$$C_{\text{cal}} = 10.84 \text{ kJ/K}$$

Thermochemistry Exercises Answers

E. N. Ramsden



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Solutions Manual to Accompany Elements of Physical Chemistry C. A. Trapp, M. P. Cady, 2007 The Solutions manual to accompany Elements of Physical Chemistry 4e contains full worked solutions to all end of chapter exercises featured in the book

Thermochemistry Mr. Ajeet Singh, 2018-05-16 Introduces thermodynamic principles related to chemical processes including enthalpy entropy Gibbs free energy and calorimetry

Ions in Solution J Burgess, 1999-10-01 This outline of the principles and chemical interactions in inorganic solution chemistry delivers a course module in an area of considerable complexity Problems with solutions and tutorial hints to test comprehension have been added as a feature to check readers understanding and assist self study Exercises and projects are also provided to help readers deepen and extend their knowledge and understanding Inorganic solution chemistry is treated thoroughly Emphasis is placed upon NMR UV VIS IR Raman spectroscopy X ray diffraction and such topics as acid base behaviour stability constants and kinetics

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Student

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K. S. Verma, 2022-05-19 Physical Chemistry for JEE Advanced Part 1 a Cengage Exam Crack Series product is designed to help aspiring engineers focus on the subject of physical chemistry from two standpoints To develop their caliber aptitude and attitude for the engineering field and profession To strengthen their grasp and understanding of the concepts of the subjects of study and their applicability at the grassroots level Each book in this series approaches the subject in a very conceptual and coherent manner While its illustrative solved examples facilitate easy mastering of the concepts and their applications an array of solved problems exposes the students to a variety of questions that they can expect in the examination The coverage and features of this series of books make it highly useful for all those preparing for JEE Main and Advanced and aspiring to become engineers

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