

Thermochemistry Formula Sheet:

1st Law of Thermodynamics: Endothermic: $q = +\cancel{L}$ Exothermic: $q = -\cancel{L}$	Internal Energy: $\Delta U = q + W$
Work Done By/On a Gas: Work done by the system: $W = -\cancel{L}$ Work done on the system: $W = +\cancel{L}$	$W = -P \Delta V$ $1 L \cdot atm = 101.3 \text{ Joules}$
Specific Heat Capacity: $J/(g \cancel{L} ^\circ C) \cancel{L}$	$C_s = \frac{q}{m \Delta T}$
Molar Heat Capacity: $J/(mol ^\circ C)$	$C_M = \frac{q}{n \Delta T}$
Heat Capacity: $J/^\circ C$	$C_n = \frac{q}{\Delta T}$
Calorimetry - Temperature Change:	$q = mc \Delta T$ $C_{H_2O} = 4.184 \frac{J}{g ^\circ C} \quad C_{ice} = 2.09 \frac{J}{g ^\circ C} \quad C_{steam} = 2.03 \frac{J}{g ^\circ C}$
Calorimetry - Phase Change:	$q = m \Delta H \quad q = n \Delta H$ $H_2O: \Delta H_{fus} = 334 J/g \quad \Delta H_{fus} = 6.01 kJ/mol$ $H_2O: \Delta H_{vap} = 2260 J/g \quad \Delta H_{vap} = 40.7 kJ/mol$
Final Temperature of a Mixture: $T_A \rightarrow$ Initial Temp. of Substance A $T_B \rightarrow$ Initial Temp. of Substance B	$-q_A = q_B$ $-m_A C_A (T_f - T_A) = m_B C_B (T_f - T_B)$ Note: This equation will work if there are no phase changes.
Coffee Cup Calorimeter:	$\Delta H^\circ_{rxn} = \frac{q_{rxn}}{n} \quad q_{rxn} = -q_{H_2O}$
Enthalpy Change:	

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