

Thermochemistry Practice Problems

KEY

Flow of Energy:

1. How many kJ of energy are in a donut that contains 200.0 Calories?

$$\frac{200 \text{ Cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ Cal}} = \boxed{836.8 \text{ kJ}}$$

2. What is the specific heat of a substance that has a mass of 25.0 g and requires 525.0 calories to raise its temperature by 15.0°C?

$$\begin{aligned} m &= 25.0 \text{ g} \\ q &= 525.0 \text{ cal} \\ \Delta T &= 15.0^\circ\text{C} \end{aligned}$$

$$q = mc\Delta T$$

$$525.0 \text{ cal} = (25.0 \text{ g})(c)(15.0^\circ\text{C})$$

$$\boxed{c = 1.40 \text{ cal/g}^\circ\text{C}}$$

3. Suppose 100.0 g of H₂O(s) absorbs 1255.0 J of heat. What is the corresponding temperature change? The specific heat capacity of H₂O(s) is 2.1 J/g°C.

$$\begin{aligned} q &= 1255.0 \text{ J} \\ m &= 100.0 \text{ g} \\ c &= 2.1 \text{ J/g}^\circ\text{C} \end{aligned}$$

$$q = mc\Delta T$$

$$1255.0 \text{ J} = (100.0 \text{ g})(2.1 \text{ J/g}^\circ\text{C})(\Delta T)$$

$$\boxed{\Delta T = 6.0^\circ\text{C}}$$

4. How many joules of heat energy are required to raise the temperature of 100.0 g of aluminum by 120.0°C? The specific heat capacity of aluminum is 0.90 J/g°C.

$$\begin{aligned} m &= 100.0 \text{ g} \\ c &= 0.90 \text{ J/g}^\circ\text{C} \\ \Delta T &= 120.0^\circ\text{C} \end{aligned}$$

$$q = (100.0 \text{ g})(0.90 \text{ J/g}^\circ\text{C})(120.0^\circ\text{C})$$

$$\boxed{q = 11,000 \text{ J}}$$

5. In reference to #4, what is the molar heat capacity of aluminum?

$$(0.90 \text{ J/g}^\circ\text{C})(26.98 \text{ g/mol}) = \boxed{24 \text{ J/mol}^\circ\text{C}}$$

Measuring and Expressing Heat Changes:

6. A student mixed 75.0 mL of water containing 0.75 mol HCl at 25°C with 75.0 mL of water containing 0.75 mol of NaOH at 25°C in a foam cup calorimeter. The temperature of the resulting solution increased to 35°C. How much heat in kilojoules was released by this reaction ($C_{\text{water}} = 4.18 \text{ J/g}^\circ\text{C}$)?

$$q = mc\Delta T$$

$$m = 150.0 \text{ g}$$

$$c = 4.184 \text{ J/g}^\circ\text{C}$$

$$\Delta T = 10^\circ\text{C}$$

$$q = (150.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(10^\circ\text{C})$$

$$\boxed{q = 6300 \text{ J}}$$

7. Calculate the amount of heat evolved when 150 g of water are reacted with excess calcium according to the following thermochemical equation:



$$\frac{150 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}} \times \frac{-65.2 \text{ kJ}}{1} = \boxed{-270 \text{ kJ}; \text{ exothermic}}$$

Thermochemistry Guided Practice Problems

Tao Wei



Thermochemistry Guided Practice Problems:

This book delves into Thermochemistry Guided Practice Problems. Thermochemistry Guided Practice Problems is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Thermochemistry Guided Practice Problems, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Thermochemistry Guided Practice Problems
 - Chapter 2: Essential Elements of Thermochemistry Guided Practice Problems
 - Chapter 3: Thermochemistry Guided Practice Problems in Everyday Life
 - Chapter 4: Thermochemistry Guided Practice Problems in Specific Contexts
 - Chapter 5: Conclusion
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