

Thermochemistry Test Review: Multiple Choice Problems

- What is the correct order, from highest to lowest molar heat value, for the three types of energy change?
 - Chemical, phase, nuclear
 - Phase, chemical, nuclear
 - Nuclear, phase, chemical
 - Nuclear, chemical, phase
- Which metal requires the most energy to raise 1.00 g of it by 1.00°C? (Refer to Chemistry Data Booklet)
 - Aluminum
 - Copper
 - Lead
 - Nickel
- What is the specific heat capacity of a substance if 2.41×10^4 J are needed to change the temperature of 105.0 g of it from 25.0°C to 250.0°C?
 - 1.02×10^{-4} J/g°C
 - 9.18×10^{-4} J/g°C
 - 0.918 J/g°C
 - 1.02 J/g°C
- Which statement about enthalpy is true?
 - Heat is given off to the surroundings in endothermic reactions.
 - Some substances have a negative specific heat capacity.
 - Specific heat capacity is the same for all liquids.
 - The sign of ΔH is always negative in exothermic reactions.
- What happens to the value of ΔH for a thermochemical reaction if the reaction is reversed?
 - ΔH has the same numerical value, and the sign changes.
 - ΔH has the same numerical value, and the sign remains the same.
 - ΔH is the reciprocal of the original value, and the sign changes.
 - ΔH is the reciprocal of the original value, and the sign remains the same.
- Which is an exothermic process?
 - Ice melting
 - Water boiling
 - Water evaporating
 - Water vapour condensing
- Which statement is true for the combustion of ethanol?

$$\text{C}_2\text{H}_5\text{OH}(\text{l}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{l}) \quad \Delta H = -1370 \text{ kJ}$$
 - The enthalpy change would be the same if gaseous water were produced.
 - The potential energy of the products is less than the potential energy of the reactants.
 - The products of the reaction occupy a larger volume than the reactants.
 - The reaction is endothermic.
- What is the energy required to evaporate two moles of liquid water given the following equations?

$$\begin{aligned} 2\text{H}_2\text{O}(\text{g}) + \text{O}_2(\text{g}) &\rightarrow 2\text{H}_2\text{O}(\text{l}) & \Delta H &= -483.6 \text{ kJ} \\ 2\text{H}_2\text{O}(\text{g}) + \text{O}_2(\text{g}) &\rightarrow 2\text{H}_2\text{O}(\text{g}) & \Delta H &= -371.6 \text{ kJ} \end{aligned}$$
 - 44.0 kJ
 - 88.0 kJ
 - 327.6 kJ
 - 1053.2 kJ
- Which statement correctly describes an endothermic chemical reaction?
 - The products have higher potential energy than the reactants, and the ΔH is negative.
 - The products have higher potential energy than the reactants, and the ΔH is positive.
 - The products have lower potential energy than the reactants, and the ΔH is negative.
 - The products have lower potential energy than the reactants, and the ΔH is positive.
- Given the thermochemical equation,

$$2\text{NO}_2(\text{g}) \rightarrow 2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \quad \Delta H = +114 \text{ kJ}$$
 what is the ΔH for the reaction $\text{NO}(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$?
 - 114 kJ
 - 57 kJ
 - +57 kJ
 - +114 kJ
- Using the data below, what is the order of changes that occur when ethanol is heated from -25.0°C to 85.0°C?

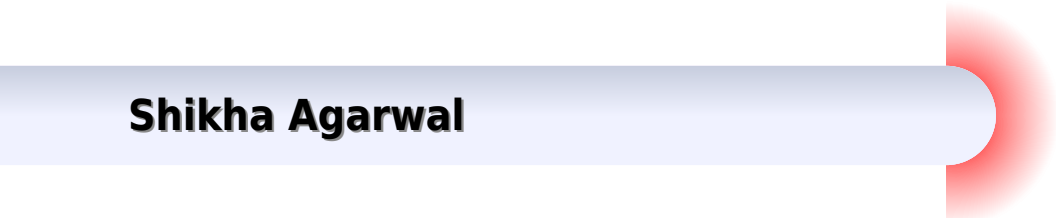
Boiling point of ethanol: 78.3°C
Melting point of ethanol: -117.3°C

 - Phase change, temperature change
 - Phase change, temperature change, phase change
 - Temperature change, phase change
 - Temperature change, phase change, temperature change
- Which statement best describes the formation of methanol, $\text{CH}_3\text{OH}(\text{l})$, from the elements? (Refer to Data Booklet, page 5)
 - It is endothermic and heat is absorbed.
 - It is exothermic and heat is released.
 - It is exothermic and heat is absorbed.
 - It is exothermic and heat is released.
- Which process results in the greatest endothermic change for 10.0 g of H_2O ?
 - Condensation
 - Melting
 - Solidification
 - Vaporization
- The following decomposition reaction may occur in an air bag:

$$2\text{NaN}_3(\text{s}) \rightarrow 3\text{N}_2(\text{g}) + 2\text{Na}(\text{s}) \quad \Delta H = -43.5 \text{ kJ}$$
 What is the heat of formation, ΔH_f° , for NaN_3 ?
 - 43.5 kJ
 - 21.8 kJ
 - 21.8 kJ
 - 43.5 kJ
- A small sample released 2.0×10^{11} kJ of energy while undergoing a change. What type of change most likely occurred?
 - Chemical
 - molecular
 - nuclear
 - physical

Thermochemistry Multiple Choice Problems

Shikha Agarwal



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Internal Combustion Engines Shyam K. Agrawal, 2006 Salient Features The New Edition Is A Thoroughly Revised Version Of The Earlier Edition And Presents A Detailed Exposition Of The Basic Principles Of Design Operation And Characteristics Of Reciprocating I C Engines And Gas Turbines Chemistry Of Combustion Engine Cooling And Lubrication Requirements Liquid And Gaseous Fuels For Ic Engines Compressors Supercharging And Exhaust Emission Its Standards And Control Thoroughly Explained Jet And Rocket Propulsion Alternate Potential Engines Including Hybrid Electric And Fuel Cell Vehicles Are Discussed In Detail Chapter On Ignition System Includes Electronic Injection Systems For Si And Ci Engines 150 Worked Out Examples Illustrate The Basic Concepts And Self Explanatory Diagrams Are Provided Throughout The Text More Than 200 Multiple Choice Questions With Answers A Good Number Of Review Questions Numerical With Answers For Practice Will Help Users In Preparing For Different Competitive Examinations With These Features The Present Text Is Going To Be An Invaluable One For Undergraduate Mechanical Engineering Students And Amie Candidates **Thermochemistry** Mr.

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