

Practice Questions: Thermochemistry

- Calculate the work energy, w , gained or lost by the system when a gas expands from 15 L to 35 L against a constant pressure of 1.5 atm. (1 L·atm = 100 J)
- What is expected when the reaction shown below takes place in a thermally-insulated container outfitted with a movable piston at a constant atmospheric pressure of 1 atm?
 $2\text{C}_2\text{H}_2(g) + 7\text{O}_2(g) \rightarrow 4\text{CO}_2(g) + 6\text{H}_2\text{O}(g)$
 - Volume will decrease and work will be done on the system.
 - Volume will decrease and work will be done by the system.
 - Volume will increase and work will be done by the system.
 - Volume will decrease and work will be done on the system.
- An ideal gas expands into a vacuum (external pressure = 0) without gaining or losing heat. For this expansion...
 - ΔE increases.
 - ΔE decreases.
 - ΔE does not change.
- Which of the following is TRUE if $E_{\text{sys}} = -95 \text{ J}$?
 - The system is gaining 95 J, while the surroundings are losing 95 J.
 - The system is losing 95 J, while the surroundings are gaining 95 J.
 - Both the system and the surroundings are gaining 95 J.
 - Both the system and the surroundings are losing 95 J.
 - None of the above is true.
- Which of the following signs for q and w represent a system that is doing work on the surroundings, as well as losing heat to the surroundings?
 - $q = -$, $w = -$
 - $q = -$, $w = +$
 - $q = +$, $w = +$
 - $q = +$, $w = -$
 - None of these represent the system referenced above.
- In which case is the work on the system always positive?
 - $\Delta V < 0$
 - $\Delta E > 0$
 - $\Delta V > 0$
 - $\Delta V = 0$
 - $\Delta E = 0$
 - $\Delta E < 0$
- What is the change in internal energy for a system that absorbs 15.0 kJ of heat and performs 610 J of work on the surroundings?
- Calculate the change in internal energy (ΔE) for a system that is giving off 25.0 kJ of heat and is changing from 12.00 L to 6.00 L in volume at 1.50 atm pressure. (Remember that 101.3 J = 1 L·atm.)
- Find ΔE° for the reaction below if the process is carried out at a constant pressure of 1.00 atm and ΔV (the volume change) equals -24.5 L . (1 L·atm = 101 J)
 $2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g) \quad \Delta H^\circ = -566 \text{ kJ}$
- Which of the following results in the largest decrease (most negative change) in internal energy for a system?
 - The system volume decreases and the surroundings get colder.
 - The system volume decreases and the surroundings get hotter.
 - The system volume increases and the surroundings get colder.
 - The system volume increases and the surroundings get hotter.
 - The system volume remains constant and the surroundings get colder.

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Neil D. Jespersen, Pamela Kerrigan



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Table of Contents Thermochemistry Practice Multiple Choice Questions

1. Understanding the eBook Thermochemistry Practice Multiple Choice Questions
 - The Rise of Digital Reading Thermochemistry Practice Multiple Choice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Practice Multiple Choice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Practice Multiple Choice Questions
 - Personalized Recommendations
 - eBook Platform User Reviews and Ratings
 - eBook Platform Bestseller Lists

5. Accessing Thermochemistry Practice Multiple Choice Questions Free and Paid eBooks
 - Thermochemistry Practice Multiple Choice Questions Public Domain eBooks
 - Thermochemistry Practice Multiple Choice Questions eBook Subscription Services
 - Thermochemistry Practice Multiple Choice Questions Budget-Friendly Options
6. Navigating Thermochemistry Practice Multiple Choice Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Practice Multiple Choice Questions Compatibility with Devices
 - Thermochemistry Practice Multiple Choice Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Practice Multiple Choice Questions
 - Highlighting and Note-Taking Thermochemistry Practice Multiple Choice Questions
 - Interactive Elements Thermochemistry Practice Multiple Choice Questions
8. Staying Engaged with Thermochemistry Practice Multiple Choice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Practice Multiple Choice Questions
9. Balancing eBooks and Physical Books Thermochemistry Practice Multiple Choice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Practice Multiple Choice Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermochemistry Practice Multiple Choice Questions
 - Setting Reading Goals Thermochemistry Practice Multiple Choice Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermochemistry Practice Multiple Choice Questions
 - Fact-Checking eBook Content of Thermochemistry Practice Multiple Choice Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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