

PRACTICE TEST/THERMOCHEMISTRY/AP CHEMISTRY

1. a) Calculate the amount of heat transferred when 10.00 g of $\text{N}_2\text{O}(\text{g})$ is formed by the following reaction:



- b) Draw an energy diagram for this process.

2. Predict the value for ΔH°_f for the following scenarios:

- a) $\text{Br}_2(\text{g})$
- b) $\text{Br}_2(\text{l})$
- c) $\text{I}_2(\text{g})$
- d) $\text{I}_2(\text{s})$

3. Calculate the ΔH_{rxn} for the following reaction:



$$\begin{aligned}\Delta H^\circ_f \text{ C}_2\text{H}_4(\text{g}) &= 226.6 \text{ kJ/mol} \\ \Delta H^\circ_f \text{ CO}_2(\text{g}) &= -393.5 \text{ kJ/mol} \\ \Delta H^\circ_f \text{ H}_2\text{O}(\text{l}) &= -285.8 \text{ kJ/mol}\end{aligned}$$

4. A 5.00 g sample of liquid water at 25.0°C is heated by the addition of 84.0 J of energy. The final temperature of the water is _____ $^\circ\text{C}$. The specific heat capacity of the liquid is $4.18 \text{ J/g}^\circ\text{C}$.

- a) -21 b) 4.02 c) 25.2 d) 95.2 e) 29.0

5. *THIS PROBLEM WAS TAKEN DIRECTLY FROM THE 1995 AP CHEM. EXAM.*
Propane is a hydrocarbon that is commonly used as a fuel for cooking.
Propane's formula is C_3H_8 .

- a) Write a balanced equation for the complete combustion of propane gas.
- b) Calculate the volume of air at 30°C and 1.00 atm that is needed to burn completely 10.0 g of propane. Assume that air is 21.0% O_2 by volume.
- c) The heat of combustion ($\Delta H^\circ_{\text{combustion}}$) is $-2,220.1 \text{ kJ/mol}$. Calculate the heat of formation, ΔH°_f , of propane given that ΔH_f of $\text{H}_2\text{O}(\text{l})$ is -285.3 kJ/mol and ΔH°_f of $\text{CO}_2(\text{g})$ is -393.5 kJ/mol .
- d) Assuming that all of the heat evolved burning 10.0 g propane is transferred to 8.00 kg of water (specific heat = $4.184 \text{ J/g}^\circ\text{C}$), calculate the increase in temperature of the water.

Thermochemistry Practice Test

Neil D. Jespersen, Pamela Kerrigan



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Table of Contents Thermochemistry Practice Test

1. Understanding the eBook Thermochemistry Practice Test
 - The Rise of Digital Reading Thermochemistry Practice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Practice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Practice Test
 - Personalized Recommendations
 - Thermochemistry Practice Test User Reviews and Ratings
 - Thermochemistry Practice Test and Bestseller Lists
5. Accessing Thermochemistry Practice Test Free and Paid eBooks
 - Thermochemistry Practice Test Public Domain eBooks
 - Thermochemistry Practice Test eBook Subscription Services

- Thermochemistry Practice Test Budget-Friendly Options
- 6. Navigating Thermochemistry Practice Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Practice Test Compatibility with Devices
 - Thermochemistry Practice Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Practice Test
 - Highlighting and Note-Taking Thermochemistry Practice Test
 - Interactive Elements Thermochemistry Practice Test
- 8. Staying Engaged with Thermochemistry Practice Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Practice Test
- 9. Balancing eBooks and Physical Books Thermochemistry Practice Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Practice Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Practice Test
 - Setting Reading Goals Thermochemistry Practice Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Practice Test
 - Fact-Checking eBook Content of Thermochemistry Practice Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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