



Unit: Thermodynamics "Thermodynamic Equations" - Week 8 10/26/2017

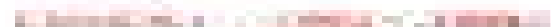


Answer these 4 questions using the following problems using dimensional analysis. Show all work.

1. The combustion enthalpy, ΔH_c , is exothermic reaction. Use the information below to determine amount of energy given off when 2.5 moles of ethanol is combusted.



Answer: 3415.25 kJ



2. Combustion heat for ethanol from the reaction of oxygen, ethane, and oxygen. How many moles of water can be made that liberates 10.0 kJ of energy?



Answer: 0.1025 moles

3. Ethanol combustion equation, $C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O$ liberates 1366.1 kJ of energy. How many moles of ethanol are required to produce 10.0 moles of water?



Answer: 3.33 moles

4. Burning 1 mole, C_2H_5OH , liberates 1366.1 kJ of energy. How many moles of H_2O are produced by 10.0 moles of burning ethanol?



Answer: 30 moles

Thermochemistry Review Answer Key Solutions

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