

Thermochemistry Review Part 1

Definition Questions: you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is $0.5050 \text{ J/g}^\circ\text{C}$. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of 50g and an initial temperature of 90°C is placed into 100mL of water at a temperature of 25°C . The temperature of water rises to 31.3°C . Determine the specific heat capacity of aluminum.

Q lost (Al) = Q gained (Water)

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3-90)) = (100 \times 4.18 \times (31.3-25))$$

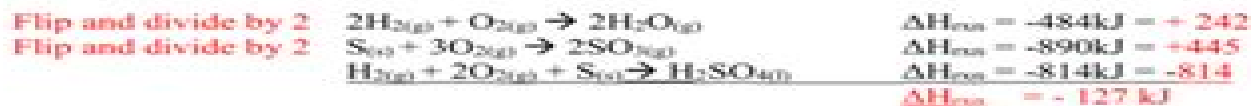
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by 300g of methanol, CH_3OH , as it evaporates? Given: The molar heat of vaporization is 35.3kJ/mol .

$q = ?$	$\Delta H = q/n$	$n \text{ CH}_3\text{OH} = m/\text{MM}$
$m = 300 \text{ g}$	$q = H \times n$	$n = 300\text{g} / 32 \text{ g/mol}$
$\Delta H = 35.3 \text{ kJ/mol}$	$q = (35.5 \text{ kJ} \times 9.375 \text{ mol})$	$n = 9.375 \text{ mol}$
	$q = 332.81 \text{ kJ}$	

4. What is the molar enthalpy of the formation of $1 \text{ mol H}_2\text{SO}_4(\text{l})$ given the following information?



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1. Understanding the eBook Thermochemistry Test Answer Key
 - The Rise of Digital Reading Thermochemistry Test Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Test Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Test Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Test Answer Key
 - Personalized Recommendations
 - Thermochemistry Test Answer Key User Reviews and Ratings
 - Thermochemistry Test Answer Key and Bestseller Lists
5. Accessing Thermochemistry Test Answer Key Free and Paid eBooks
 - Thermochemistry Test Answer Key Public Domain eBooks
 - Thermochemistry Test Answer Key eBook Subscription Services
 - Thermochemistry Test Answer Key Budget-Friendly Options

6. Navigating Thermochemistry Test Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Test Answer Key Compatibility with Devices
 - Thermochemistry Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Test Answer Key
 - Highlighting and Note-Taking Thermochemistry Test Answer Key
 - Interactive Elements Thermochemistry Test Answer Key
8. Staying Engaged with Thermochemistry Test Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Test Answer Key
9. Balancing eBooks and Physical Books Thermochemistry Test Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Test Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermochemistry Test Answer Key
 - Setting Reading Goals Thermochemistry Test Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermochemistry Test Answer Key
 - Fact-Checking eBook Content of Thermochemistry Test Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

-
- Interactive and Gamified eBooks

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