Experiment 4: Standardization of Sodium Thiosulfate Solution

Introduction

In the first experiment, we performed a gravimetric analysis experiment by measuring the mass of a precipitate. In this experiment, we will be performing a volumetric analysis in which we will determine the amount of a constituent by measuring the volume of a reagent. In a volumetric analysis, it is necessary to have a solution of known concentration (also known as standard solution) reacting with a sample of unknown concentration to determine the unknown concentration. Such a technique is called titration.

A standard solution may be prepared by the simple procedure of taking a known amount of solid and diluting it to make a definite volume of the solution; in this case, the term primary standard solution is applied. Alternatively, it is very often the case that a direct measurement of the amount of material cannot be made, or is very difficult to make accurately. Such examples could be if the sample is a gas, a volatile liquid, a solid that is hygroscopic (has a tendency to absorb water and dissolve) or a solid that is efflorescent (has a tendency to lose water from its molecular structure). In this case, a solution must be made up with approximately the required concentration and the actual concentration must be determined by titration with a primary standard solution. The term secondary standard solution is applied in this latter instance. Hence:

- A primary standard solution is prepared by direct measurements of the mass of solute and the volume of solution.
- A secondary standard solution is a solution whose concentration cannot be determined directly from the weight of solute and volume of solution; the concentration must be determined by analysis of the solution itself.

Sodium thiosulfate crystals have the formula Na₂S₂O₃•5H₂O when freshly crystallized, but the salt tends to effloresce (dehydrate) some of its water of crystallization (a property that is increased in very dry areas). Since the true formula for the crystals changes over time, we cannot weigh the sample and divide by the molar mass to determine moles present. This is why sodium thiosulfate cannot be used as a primary standard. A solution of thiosulfate is prepared with approximately the desired concentration and then standardized with a primary standard solution such as potassium iodate. The process of determining the concentration of the secondary standard solution by titration is called 'standardization'.

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