

DATE: 10/20/21
 SECTION/GROUP: #10
 NAME: Bailey Spell
 NAME: Gavin Sprinkle
 NAME: Ben Steadley

Investigation 3 Written Argument

Guiding Question: What is the concentration of acetic acid in vinegar?

Claim: My group and I claim that the concentration of acetic acid in vinegar is 0.83 M, 4.99% m/v

Evidence: Balanced equation: $1\text{CH}_3\text{COOH(aq)} + 1\text{NaOH(aq)} \gg> 1\text{NaCH}_3\text{COO(aq)} + 1\text{H}_2\text{O(l)}$

Data Table

Data:	Trial 1:	Trial 2:	Trial 3:	Trial 4:
Molarity of NaOH used	0.400 M	0.400 M	0.400 M	0.400 M
Volume of Acetic Acid used	10.00 mL	10.00 mL	10.00 mL	10.00 mL
Initial Burette Reading (NaOH)	0.35 mL	21.15 mL	0.49 mL	21.21 mL
Final Burette Reading (NaOH)	21.15 mL	41.87 mL	21.21 mL	42.15 mL
Volume of NaOH used	20.80 mL	20.72 mL	20.72 mL	20.94 mL
Shade of Pink	Light pink	Light pink	Dark pink	Dark pink

Calculations of volume of NaOH used:

Trial 1: Initial Reading - Final Reading

21.15 mL NaOH - 0.35 mL of NaOH = 20.80 mL NaOH

Trial 2: Initial Reading - Final Reading

41.87 mL NaOH - 21.15 mL of NaOH = 20.72 mL NaOH

Trial 3: Initial Reading - Final Reading

21.21 mL NaOH - 0.49 mL NaOH = 20.72 mL NaOH

Trial 4: Initial Reading - Final Reading

42.15 mL NaOH - 21.21 mL NaOH = 20.94 mL NaOH

Calculations of moles of NaOH used:

Moles of NaOH used = Molarity x Volume in L

Make sure to change mL of NaOH into L of NaOH

Trial 1: $20.80 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02080 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02080 \text{ L NaOH} = 0.00832 \text{ mols NaOH}$

Trial 2: $20.72 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02072 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02072 \text{ L NaOH} = 0.00829 \text{ mols NaOH}$

Trial 3: $20.72 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02072 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02072 \text{ L NaOH} = 0.00829 \text{ mols NaOH}$

Trial 4: $20.94 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02094 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02094 \text{ L NaOH} = 0.00838 \text{ mols NaOH}$

Calculations of moles and Molarity of Acetic Acid:

mols of Acetic Acid is same as mols of NaOH due to balanced equation

Make sure to change mL of volume of acetic acid used to L of acetic acid used $\gg> \gg$

$(10.00 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}}) / 1000 \text{ mL} = 0.010 \text{ L}$

Molarity = mols/L

Trial 1: $0.00832 \text{ mols A.A.} / 0.010 \text{ L A.A.} = 0.83 \text{ M A.A.}$

Trial 2: $0.00829 \text{ mols A.A.} / 0.010 \text{ L A.A.} = 0.83 \text{ M A.A.}$

Us Ws 1 V3 0 Answers

Michael Brown

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

Us Ws 1 V3 0 Answers:

Us Ws 1 V3 0 Answers: Bestsellers in 2023 The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous captivating novels captivating the hearts of readers worldwide. Lets delve into the realm of bestselling books, exploring the engaging narratives that have enthralled audiences this year. Us Ws 1 V3 0 Answers : Colleen Hoover "It Ends with Us" This touching tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover masterfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can succeed. Uncover the Best : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This captivating historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids captivating storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Us Ws 1 V3 0 Answers : Delia Owens "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens weaves a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These popular novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a masterful and thrilling novel that will keep you speculating until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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