

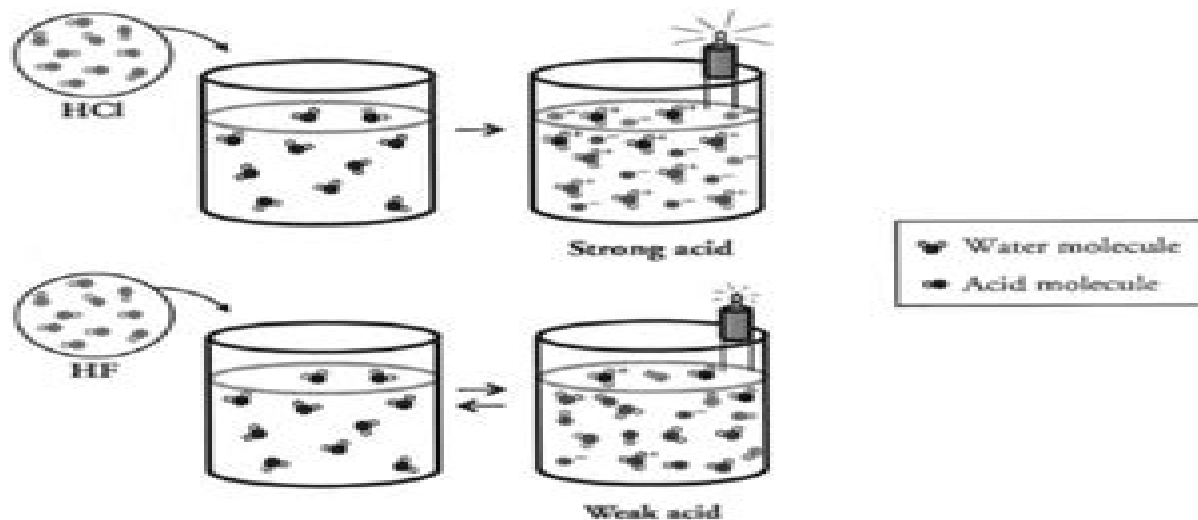
Strong versus Weak Acids

What makes a strong acid strong?

Why?

Acids are substances that surround us in our everyday life. The uses of acids range from providing essential nutrients for our bodies to dissolving metals. Some acids are safe to handle with our bare hands or even use in food preparation. Other acids will severely burn human skin. It is important to understand how these substances can all be acids and yet have such different properties.

Model 1 – Acid Strength and Conductivity



Note: Excess water molecules in the solutions are not shown.

1. Examine the symbols in Model 1. Match each symbol with its correct meaning.

- | | | | | | |
|----------|---------------|----------|--------------------|----|--------------------|
| <u>A</u> | water | <u>B</u> | acid | a. | Water molecule |
| <u>d</u> | Hydronium ion | <u>c</u> | conjugate base ion | b. | Acid molecule |
| | | | | c. | Conjugate base ion |
| | | | | d. | Hydronium ion |

Strong Versus Weak Acids Answer Key

Victor M. Corman



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