

Name \_\_\_\_\_ Date \_\_\_\_\_

## Fruit Fly Lab

Humans have 46 chromosomes in 23 homologous pairs. By comparison, fruit flies have 8 chromosomes in 4 homologous pairs. Fewer chromosomes make fruit flies good candidates for analyzing the inheritance of traits. [1]

Rather than raising fruit flies this lab will use the simulator found at [FlyLab.UC](http://FlyLab.UC). The menu on the left side of the screen provided traits that can be altered. Pictures of the flies being crossed are shown in the center of the screen. Play with the software for a minute or two to familiarize yourself. Complete the actions for each of the following scenarios.

### Section 1: Setting up a cross

1. Upon opening the simulation a wild type fly is pictured. Wild type traits are the traits most commonly found in nature. Traits other than the wild type are the result of mutation and are not as common. Click the magnifying glass in the bottom right of the screen to get a better look at the P1 wild type female fly.
2. Click the [Select Fly for Mating] box.
3. The [Flies to be Mated] screen will open. Beneath the shadow outline of the male click the [Design Fly] box followed by the [Select fly for mating] box.
4. Click the [Mate Flies] box.
5. Open the **RESULTS** tab along the top of the screen. Use the results to fill in the following information. Note: the + sign means that the traits are wild-type or normal.

| Phenotype | Observed | Proportion |
|-----------|----------|------------|
|           |          |            |
|           |          |            |
| Total     |          |            |

6. What is the only difference between the offspring for this cross?
7. How is the proportion calculated?

# Virtual Fruit Fly Lab Answers

**Leigh Ann Henion**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta H = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta H = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta H = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_3\text{O}^+$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = -55.8 \text{ kJ}$  as shown note

that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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