

Atomic Theory Worksheet

Name: _____

John Dalton (1766 – 1844):

John Dalton was an English chemist. His ideas form the atomic theory of matter. Here are his ideas.

Atomic Theory



- All elements are composed (made up) of atoms. It is impossible to divide or destroy an atom.
- All atoms of the same elements are alike. (One atom of oxygen is like another atom of oxygen.)
- Atoms of different elements are different. (An atom of oxygen is different from an atom of hydrogen.)
- Atoms of different elements combine to form a compound. These atoms must be in definite whole number ratios. For example, water is a compound made up of 2 atoms of hydrogen and 1 atom of oxygen (a ratio of 2:1). Three atoms of hydrogen and 2 atoms of oxygen cannot combine to make water.



1803

1. What is the name of John Dalton's theory? _____
2. What are elements made of? _____
3. An atom of hydrogen and an atom of carbon are _____.
4. What are compounds made of? _____
5. The ratio of atoms in HCl is: a) 1:3 b) 2:1 c) 1:1

J. J. Thompson (Late 1800s):

J. J. Thompson was an English scientist. He discovered the electron when he was experimenting with gas discharge tubes. He noticed a movement in a tube. He called the movement cathode rays. The rays moved from the negative end of the tube to the positive end. He realized that the rays were made of negatively charged particles – electrons.

Atomic Structure Model



J. J. Thompson



1904

1. What did J.J. Thompson discover? _____
2. What is the charge of an electron? _____
3. What are cathode rays made of? _____
4. Why do electrons move from the negative end of the tube to the positive end?

5. What was Thompson working with when he discovered the cathode rays?

NUCLEAR MODEL



ERNEST RUTHERFORD



1911

Lord Ernest Rutherford (1871 – 1937):

Ernest Rutherford conducted a famous experiment called the gold foil experiment. He used a thin sheet of gold foil. He also used special equipment to shoot alpha particles (positively charged particles) at the gold foil. Most particles passed straight through the foil like the foil was not there. Some particles went straight back or were deflected (went in another direction) as if they had hit something. The experiment shows:

- Atoms are made of a small positive nucleus; positive nucleus repels (pushes away) positive alpha particles
- Atoms are mostly empty space

1. What is the charge of an alpha particle? _____
2. How did he know that an atom was mostly empty space?

3. What happened to the alpha particles as they hit the gold foil?

4. How did he know that the nucleus was positively charged?

Student 1 Answers Atomic Theory

Franklin Bache



Student 1 Answers Atomic Theory:

Exploration of the Structure of Atom Chandan Sengupta, First Publication April 2019 Revised Publication October 2022 Second Revised Edition July 2023 Third Revised Edition December 2024 Published by Chandan Sengupta Printed by IECIT Printing and Publication Services Resource Centre Arabinda Nagar Bankura 722101 W B India Topics Covered Atoms and Molecules Structure of Atom Bonding Mechanism and Chemical Reactions Mechanism of Bonding This book is suitable for students of Class 9 to Class 11 Students aspiring for Pre Medical Entrance Examination can also get adequate support Additional Hard Copies can also be obtained from Chandan Sukumar Sengupta Arabinda Nagar Bankura 722101 WB Write to Us for more materials Chemistry for Degree Students B.Sc. (Honours) Semester I Madan R.L., 2022 This textbook has been designed to meet the needs of B Sc Honours First Semester students of Chemistry as per the UGC Choice Based Credit System CBCS Maintaining the traditional approach to the subject this textbook lucidly explains the basics of Inorganic and Physical Chemistry Important topics such as atomic structure periodicity of elements chemical bonding and oxidation reduction reactions gaseous state liquid state solid state and ionic equilibrium are aptly discussed to give an overview of inorganic and physical chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures **Student's Guide to Fundamentals of Chemistry** Jo A. Beran, 2013-10-22 Student's Guide to Fundamentals of Chemistry Fourth Edition provides an introduction to the basic chemical principles This book deals with various approaches to chemical principles and problem solving in chemistry Organized into 25 chapters this edition begins with an overview of how to define and recognize the more common names and symbols in chemistry This text then discusses the historical development of the concept of atom as well as the historical determination of atomic weights for the elements Other chapters consider how to calculate the molecular weight of a compound from its formula This book discusses as well the characteristics of a photon in terms of its particle like properties and defines the wavelength frequency and speed of light The final chapter deals with the fundamental components of air and the classification of materials formed in natural waters This book is a valuable resource for chemistry students lecturers and instructors *The Official ACT Prep Guide 2023-2024* ACT, 2023-05-02 The comprehensive guide to the 2023 2024 ACT test including 8 genuine full length practice tests The Official ACT Prep Guide 2023 2024 book includes six authentic ACT tests all of which contain the optional writing test so you get maximum practice before your test date These full length practice tests are also available on the Wiley Online Platform and mobile app alongside two additional bonus tests via the PIN code inside so you can study your official materials anytime anywhere This guide provides clear explanations for every answer straight from the makers of the ACT to help you improve your understanding of each subject You'll get Practical tips and strategies for boosting your score on the English math reading science and optional writing tests Eight total practice tests six in the book eight online 400 online flashcards to ensure you're mastering key concepts A customizable online test bank Expert

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A System of Chemistry for the Use of Students of Medicine

Franklin Bache, 1819

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