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### Boyle's Law

- The Volume of a fixed mass of gas varies indirectly with the pressure at constant temperature
- $PV = k$
- $P_1V_1 = P_2V_2$

### Kinetic Theory and Boyle's Law

- Pressure of a gas is caused by the molecules of the gas striking the walls of the container
- If the gas is compressed to half the volume it had, twice as many molecules are present in any given volume
  - Twice as many impacts per second on the walls of the container
  - pressure of the gas will double

Ex 1: A balloon filled with Helium has a volume of 457 mL at standard atmospheric pressure. After the balloon is released, it reaches an altitude of 6.3 km where the pressure is only 65.5 kPa. What is the volume of the balloon at this altitude?

$$P_1V_1 = P_2V_2$$

$$P \downarrow \frac{1}{2} \quad V \uparrow \times 2$$

$$101.3 \text{ kPa} \times 457 \text{ mL} = 65.5 \text{ kPa} \times V_2$$

$$V_2 = 706.8 \text{ mL}$$

Ex 2: Under a pressure of 673 mm Hg, a confined gas has a volume of 335 mL. If the pressure is increased until the volume is 252 mL, what is the new pressure, assuming the temperature remains constant?

$$P_1V_1 = P_2V_2$$

$$673 \text{ mmHg} \times 335 \text{ mL} = P_2 \times 252 \text{ mL}$$

$$P_2 = 828.2 \text{ mmHg}$$

# World Of Chemistry Episode 7 Answer Key

**Thomas Griffiths**



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### **Table of Contents World Of Chemistry Episode 7 Answer Key**

1. Understanding the eBook World Of Chemistry Episode 7 Answer Key
  - The Rise of Digital Reading World Of Chemistry Episode 7 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying World Of Chemistry Episode 7 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an World Of Chemistry Episode 7 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from World Of Chemistry Episode 7 Answer Key
  - Personalized Recommendations

- World Of Chemistry Episode 7 Answer Key User Reviews and Ratings
- World Of Chemistry Episode 7 Answer Key and Bestseller Lists
- 5. Accessing World Of Chemistry Episode 7 Answer Key Free and Paid eBooks
  - World Of Chemistry Episode 7 Answer Key Public Domain eBooks
  - World Of Chemistry Episode 7 Answer Key eBook Subscription Services
  - World Of Chemistry Episode 7 Answer Key Budget-Friendly Options
- 6. Navigating World Of Chemistry Episode 7 Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - World Of Chemistry Episode 7 Answer Key Compatibility with Devices
  - World Of Chemistry Episode 7 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of World Of Chemistry Episode 7 Answer Key
  - Highlighting and Note-Taking World Of Chemistry Episode 7 Answer Key
  - Interactive Elements World Of Chemistry Episode 7 Answer Key
- 8. Staying Engaged with World Of Chemistry Episode 7 Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers World Of Chemistry Episode 7 Answer Key
- 9. Balancing eBooks and Physical Books World Of Chemistry Episode 7 Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection World Of Chemistry Episode 7 Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine World Of Chemistry Episode 7 Answer Key
  - Setting Reading Goals World Of Chemistry Episode 7 Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of World Of Chemistry Episode 7 Answer Key
  - Fact-Checking eBook Content of World Of Chemistry Episode 7 Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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