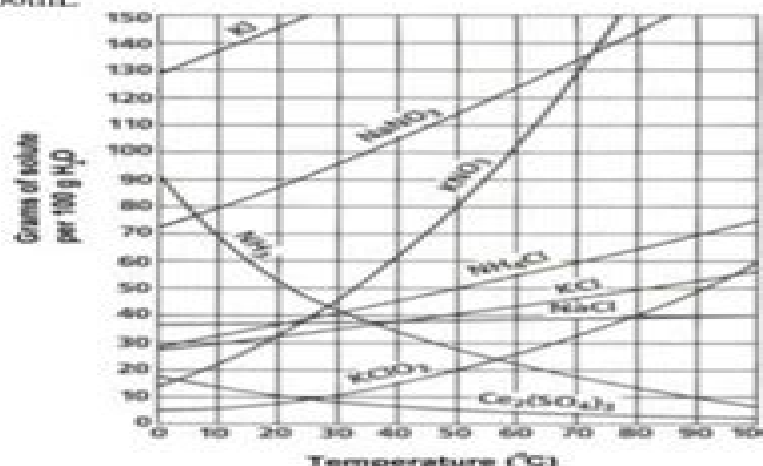


Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

### Solubility Curve Practice Problems

**Directions:** Find the mass of solute will dissolve in 100mL of water at the following temperatures?

1.  $\text{KNO}_3$  at  $70^\circ\text{C}$  = \_\_\_\_\_
2.  $\text{NaCl}$  at  $100^\circ\text{C}$  = \_\_\_\_\_
3.  $\text{NH}_4\text{Cl}$  at  $90^\circ\text{C}$  = \_\_\_\_\_
4. Which of the **above** three substances is most soluble in water at  $15^\circ\text{C}$ . = \_\_\_\_\_
5. Which of the **above** three substances is least soluble in water at  $15^\circ\text{C}$ . = \_\_\_\_\_



| 2. Solution                                                                                                    | Saturated, super saturated or Unsaturated? | If unsaturated: How much more solute can dissolve in the solution? |
|----------------------------------------------------------------------------------------------------------------|--------------------------------------------|--------------------------------------------------------------------|
| a solution that contains 70g of $\text{NaNO}_3$ at $30^\circ\text{C}$ (in 100 mL $\text{H}_2\text{O}$ )        |                                            |                                                                    |
| a solution that contains 50g of $\text{NH}_4\text{Cl}$ at $50^\circ\text{C}$ (in 100 mL $\text{H}_2\text{O}$ ) |                                            |                                                                    |
| a solution that contains 20g of $\text{KClO}_3$ at $50^\circ\text{C}$ (in 50 mL $\text{H}_2\text{O}$ )         |                                            |                                                                    |
| a solution that contains 70g of $\text{KI}$ at $0^\circ\text{C}$ (in 200 mL $\text{H}_2\text{O}$ )             |                                            |                                                                    |

3. a. How many grams of  $\text{KCl}$  can dissolve at  $5^\circ\text{C}$ ? \_\_\_\_\_  
 b. How many grams of  $\text{KCl}$  can dissolve at  $25^\circ\text{C}$ ? \_\_\_\_\_  
 c. How many grams of  $\text{Ce}_2(\text{SO}_4)_3$  can dissolve at  $10^\circ\text{C}$ ? \_\_\_\_\_  
 d. How many grams of  $\text{Ce}_2(\text{SO}_4)_3$  can dissolve at  $50^\circ\text{C}$ ? \_\_\_\_\_
4. At  $90^\circ\text{C}$ , you dissolved 10 g of  $\text{KCl}$  in 100. g of water. Is this solution saturated or unsaturated?
5. A mass of 100 g of  $\text{NaNO}_3$  is dissolved in 100 g of water at  $80^\circ\text{C}$ .  
 a) Is the solution saturated or unsaturated? \_\_\_\_\_  
 b) As the solution is cooled, at what temperature should solid first appear in the solution? Explain.
6. Use the graph to answer the following two questions:  
 Which compound is most soluble at  $20^\circ\text{C}$ ?  
 Which is the least soluble at  $40^\circ\text{C}$ ?  
 Which substance on the graph is **least** soluble at  $10^\circ\text{C}$ ?

# Unit 12 Solutions Solubility Curves Answer Key

**SA Dillow**



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